

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

- **Crystalline Solids:** Atoms, ions, or molecules are arranged in a regular, repeating three-dimensional pattern. They have sharp melting points and distinct X-ray diffraction patterns.
- **Amorphous Solids:** Atoms, ions, or molecules are arranged in a disordered, non-repeating pattern. They do not have sharp melting points and do not show distinct X-ray diffraction patterns.
- **Intermolecular Forces:** The forces between particles in a solid determine its properties. Stronger forces lead to higher melting points and greater hardness.
- **Types of Solids:**
 - Covalent Network Solids:** Atoms are bonded together in a continuous network of covalent bonds. Examples include diamond, silicon, and quartz.
 - Ionic Solids:** Consist of alternating positive and negative ions held together by strong electrostatic forces. Examples include NaCl and CaF₂.
 - Molecular Solids:** Consist of discrete molecules held together by intermolecular forces. Examples include ice, sugar, and iodine.
 - Metallic Solids:** Consist of metal cations in a sea of delocalized electrons. Examples include copper, iron, and aluminum.

Types of Chemical Calculations Allowed on the Exam

- **Stoichiometry:** Calculations involving the relative amounts of reactants and products in a chemical reaction.
- **Thermodynamics:** Calculations involving heat, work, and energy changes in a process.
- **Equilibrium:** Calculations involving the concentrations of reactants and products at equilibrium.
- **Acid-Base:** Calculations involving pH, pKa, and the concentrations of acids and bases.
- **Solubility:** Calculations involving the concentrations of ions in a saturated solution.
- **Electrochemistry:** Calculations involving cell potentials, Gibbs free energy, and the amount of substance produced or consumed in an electrochemical cell.

Periodic Table & Trends

The periodic table is color-coded to show trends. Electronegativity increases from bottom-left to top-right. Ionization energy increases from bottom-left to top-right. Atomic radius increases from top-left to bottom-right.

- **Electronegativity:** The ability of an atom to attract electrons in a chemical bond. It increases across a period and decreases down a group.
- **Ionization Energy:** The energy required to remove an electron from a neutral atom in its ground state. It increases across a period and decreases down a group.
- **Atomic Radius:** The distance from the nucleus to the outer edge of the electron cloud. It increases down a group and across a period.

Electronegativity Table

Electronegativity Table (Pauling Scale)

1.0	1.5	2.0	2.5	3.0	3.5	4.0	4.5	5.0	5.5	6.0	6.5	7.0	7.5	8.0	8.5	9.0	9.5	10.0
1.1	1.6	2.1	2.6	3.1	3.6	4.1	4.6	5.1	5.6	6.1	6.6	7.1	7.6	8.1	8.6	9.1	9.6	10.1
1.2	1.7	2.2	2.7	3.2	3.7	4.2	4.7	5.2	5.7	6.2	6.7	7.2	7.7	8.2	8.7	9.2	9.7	10.2
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1.4	1.9	2.4	2.9	3.4	3.9	4.4	4.9	5.4	5.9	6.4	6.9	7.4	7.9	8.4	8.9	9.4	9.9	10.4
1.5	2.0	2.5	3.0	3.5	4.0	4.5	5.0	5.5	6.0	6.5	7.0	7.5	8.0	8.5	9.0	9.5	10.0	10.5
1.6	2.1	2.6	3.1	3.6	4.1	4.6	5.1	5.6	6.1	6.6	7.1	7.6	8.1	8.6	9.1	9.6	10.1	10.6
1.7	2.2	2.7	3.2	3.7	4.2	4.7	5.2	5.7	6.2	6.7	7.2	7.7	8.2	8.7	9.2	9.7	10.2	10.7
1.8	2.3	2.8	3.3	3.8	4.3	4.8	5.3	5.8	6.3	6.8	7.3	7.8	8.3	8.8	9.3	9.8	10.3	10.8
1.9	2.4	2.9	3.4	3.9	4.4	4.9	5.4	5.9	6.4	6.9	7.4	7.9	8.4	8.9	9.4	9.9	10.4	10.9
2.0	2.5	3.0	3.5	4.0	4.5	5.0	5.5	6.0	6.5	7.0	7.5	8.0	8.5	9.0	9.5	10.0	10.5	11.0

Thermodynamics

First Law of Thermodynamics: Energy cannot be created or destroyed, only transferred or converted from one form to another.

Second Law of Thermodynamics: The total entropy of an isolated system can never decrease over time.

Third Law of Thermodynamics: The entropy of a perfect crystal is zero at absolute zero.

Equilibrium & Solubility

Le Chatelier's Principle: If a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will shift to counteract the change.

K_{sp} (Solubility Product Constant): A measure of the solubility of a compound. It is the product of the concentrations of the ions in a saturated solution.

Compound	K _{sp}
AgCl	1.8 × 10 ⁻¹⁰
PbCl ₂	1.7 × 10 ⁻⁵
CaF ₂	3.9 × 10 ⁻¹¹
Fe(OH) ₃	4 × 10 ⁻³⁸
Al(OH) ₃	3 × 10 ⁻³⁴
Cu(OH) ₂	2.2 × 10 ⁻²⁰
Mg(OH) ₂	1.2 × 10 ⁻¹⁷
Sn(OH) ₄	1 × 10 ⁻⁴⁰
Bi(OH) ₃	4 × 10 ⁻⁴⁰
Cr(OH) ₃	6.3 × 10 ⁻³¹
Co(OH) ₂	1.6 × 10 ⁻¹⁶
Cr(OH) ₃	6.3 × 10 ⁻³¹
Fe(OH) ₃	4 × 10 ⁻³⁸
Mn(OH) ₂	2 × 10 ⁻¹³
Ni(OH) ₂	2 × 10 ⁻¹⁶
Pb(OH) ₂	1.2 × 10 ⁻¹⁵
Sr(OH) ₂	3.2 × 10 ⁻⁴
Sn(OH) ₄	1 × 10 ⁻⁴⁰
Th(OH) ₄	1 × 10 ⁻⁴⁰
U(OH) ₄	1 × 10 ⁻⁴⁰
V(OH) ₃	5 × 10 ⁻³¹
Zn(OH) ₂	3 × 10 ⁻¹⁶

Common Ion Effect: The solubility of a salt is decreased when a common ion is added to the solution.

Buffer Solutions: Solutions that resist changes in pH when small amounts of acid or base are added. They consist of a weak acid and its conjugate base.



Acids & Bases

Arrhenius Definition: An acid is a substance that increases the concentration of H⁺ ions in aqueous solution. A base is a substance that increases the concentration of OH⁻ ions in aqueous solution.

Brønsted-Lowry Definition: An acid is a proton donor. A base is a proton acceptor.

Lewis Definition: An acid is an electron pair acceptor. A base is an electron pair donor.

Redox Reactions

Oxidation: The loss of electrons or an increase in the oxidation state of an element.

Reduction: The gain of electrons or a decrease in the oxidation state of an element.

Standard Reduction Potentials: A measure of the tendency of a chemical species to be reduced. They are listed in a table of standard reduction potentials.

Cell Potential: The potential difference between two half-cells in an electrochemical cell. It is calculated as E_{cell} = E_{cathode} - E_{anode}.

Gibbs Free Energy: A measure of the maximum work that can be done by a system at constant temperature and pressure. It is related to the cell potential by the equation ΔG = -nFE.

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Guillaume Favre

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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The Top Books of the Year Ap Chemistry Quick Study Academic The year 2023 has witnessed a noteworthy surge in literary brilliance, with numerous compelling novels captivating the hearts of readers worldwide. Lets delve into the realm of popular books, exploring the fascinating narratives that have enthralled audiences this year. Ap Chemistry Quick Study Academic :

Colleen Hoover's "It Ends with Us" This poignant tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover skillfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can succeed.

Uncover the Best : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This captivating historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids compelling storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery.

Discover the Magic : Delia Owens "Where the Crawdads Sing" This evocative coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting.

These popular novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of captivating stories waiting to be discovered.

The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts.

The Secret History is a brilliant and thrilling novel that will keep you speculating until the very end. The novel is a warning tale about the dangers of obsession and the power of evil.

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