

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

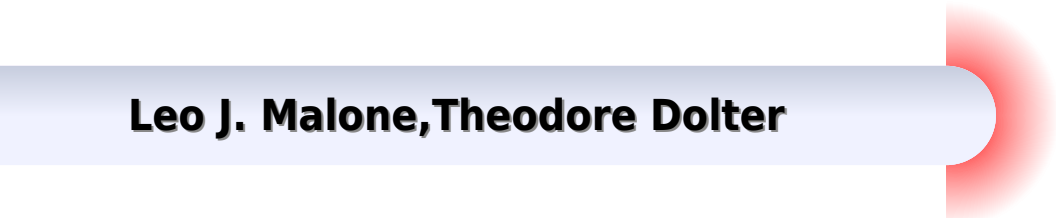
1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

Basic Concepts Of Chemistry Student Study Guide

Leo J. Malone, Theodore Dolter



Basic Concepts Of Chemistry Student Study Guide:

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Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

Organic Chemistry Study Guide with Solutions Manual Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the

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