

Chemistry Formulas



Dilution

$$M_1V_1 = M_2V_2$$

Energy of a photon

$$E_{\text{photon}} = hf = hc/\lambda$$

Pressure

$$P = F/A$$

Average Kinetic Energy

$$KE_{\text{avg}} = \frac{3}{2}RT$$

Ideal Gas Law

$$PV = nRT$$

Combined Gas Law

$$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

Dalton's Law of Partial Pressures

$$P_{\text{total}} = P_A + P_B + \dots$$

Dalton's Law of Partial Pressures

$$P_A = \chi_A P_{\text{total}}$$

Graham's Law of Effusion

$$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$$

Henry's Law

$$P_A = k_H[A]$$

Freezing Point Depression

$$\Delta T_f = -iK_f M$$

Boiling Point Elevation

$$\Delta T_b = iK_b m$$

Vapor Pressure Depression (Raoult's Law)

$$P_{\text{soln}} = \chi_{\text{soln}} P_{\text{soln}}^0$$

Osmotic Pressure (π)

$$\pi = iMRT$$

Arrhenius Equation

$$k = Ae^{-E_a/RT}$$

Equilibrium Constant Expressions

$$K_c = \frac{[\text{products}]}{[\text{reactants}]}$$

Equilibrium Constant Expressions

$$K_{eq} = \frac{k_{\text{forward}}}{k_{\text{reverse}}}$$

Equilibrium Constant Expressions

$$K_p = \frac{P_{\text{products}}}{P_{\text{reactants}}}$$

Reaction Quotient (Q)

$$Q = \frac{[\text{products}]}{[\text{reactants}]}$$

Solubility Product Constant (K_{sp})

$$K_{sp} = \frac{[\text{products}]}{[\text{reactants}]}$$

pH & pOH

$$pH = -\log[H^+]$$

pH & pOH

$$pOH = -\log[OH^-]$$

pH & pOH

$$pH + pOH = 14$$

pK_a & pK_b

pK_a & pK_b

pK_a & pK_b

Neutralization Reaction

Buffers

Enthalpy of Formation

First Law of Thermodynamics

Pressure Volume Work

Calorimetry Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Entropy (S)

Bond Dissociation Energy

Gibb's Free Energy (ΔG)

Gibb's Free Energy (ΔG)

Gibb's Free Energy (ΔG)

Standard Cell Potential

Standard Cell Potential

Nernst Equation

Kinetics (always 1st order)

Kinetics (always 1st order)

Kinetics (always 1st order)

Nuclear Binding Energy

$$pK_a = -\log[K_a]$$

$$pK_b = -\log[K_b]$$

$$pK_a + pK_b = 14$$

$$n_A M_A V_A = n_B M_B V_B$$

$$pH = pK_a + \log[A^-]/[HA]$$

$$\Delta H_f = \sum \Delta H_f^{\circ}(\text{product}) - \sum \Delta H_f^{\circ}(\text{reactants})$$

$$\Delta E = q + w$$

$$w = -P\Delta V$$

$$q = -C_{\text{calorimeter}}\Delta T$$

$$q = mc\Delta T$$

$$q = m\Delta H_{\text{fusion}}$$

$$q = m\Delta H_{\text{vaporization}}$$

$$\Delta S = \ln S_{\text{products}} - \ln S_{\text{reactants}}$$

$$\Delta H = \sum \Delta H_{\text{reactants}} - \sum \Delta H_{\text{products}}$$

$$\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$$

$$\Delta G = \Delta G^{\circ} + RT\ln Q$$

$$\Delta G^{\circ} = -RT^{\circ}\ln(K_{eq})$$

$$E^{\circ} = E^{\circ}_{\text{reduction}} + E^{\circ}_{\text{oxidation}}$$

$$E^{\circ} = E^{\circ}_{\text{cathode}} + E^{\circ}_{\text{anode}}$$

$$E_{\text{cell}} = E^{\circ} - \frac{0.0592}{n} \log Q$$

$$N = N_0 e^{-kt}$$

$$\ln N = \ln N_0 - kt$$

$$t_{1/2} = \frac{0.693}{k}$$

$$E = \Delta mc^2$$

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KJ Lindholm-Leary



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