

FIGURE 6.1
Burning Carbon in the Atomic Level
Illustration by Thomson



1. The chemical reaction starts with just one reactant: a carbon atom and a diatomic oxygen molecule.
2. The oxygen molecule splits apart to make 2 independent oxygen atoms.
3. The carbon then joins with both oxygen atoms, leaving just one product: a carbon dioxide molecule.

This does it here: a chemical reaction proceeds. The molecules must decompose into their constituent atoms, and then the atoms recombine to form the products.

Now we can make the carbon-oxygen equation a little more meaningful by adding some information about the phase of each substance in the equation. For example, when we burn coal, the carbon is in its solid state, the oxygen is in its gaseous state, and the carbon dioxide product is also in its gaseous state. We can add these facts to the equation by putting symbols as a prelude next to each substance:



Whenever you see a small letter enclosed with parentheses in a chemical equation, you know that it is a symbol telling you the phase of the substance. Table 6.1 lists the phase abbreviations used in chemical equations. Chemists do not always use phase symbols in their chemical equations. Sometimes they are there; sometimes they are not. It is always preferable to include them because they provide important information. Don't be confused if you do not see them, though. You need to memorize the abbreviations for the phases in table 6.1.

TABLE 6.1
Phase Symbols

Phase	Symbol
Solid	(s)
Liquid	(l)
Gas	(g)
Aqueous (dissolved in water)	(aq)

It also were as complicated as things get, chemistry would be a breeze. However, things get a little tougher from here. Let's look at another important chemical reaction: the burning of methane (CH₄). Methane is the principal component of natural gas. If you have a gas stove, you want to heat, or gas furnace, this chemical reaction occurs quite frequently in your home. Once again, the combustion (burning) of methane releases quite a bit of energy. This energy

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