

ANSWER KEY: BUILD AN ATOM

PART I: ATOM SCREEN

Build an Atom simulation (<http://phet.colorado.edu/en/simulation/build-an-atom>)

1. Explore the *Build an Atom* simulation with your group. As you explore, talk about what you find.
2.
 - a) List two things your group observed in the simulation.
Responses will vary, but here is what you should see....
 - Adding protons changes the identity of the atom - adding one moves from right to left on the Periodic Table.
 - You can get the atom to be stable/unstable by adjusting the number of neutrons
 - Adding electrons will change it from a +ion to neutral atom, to -ion depending on how many electrons you add
 - Adding electrons - you can only add 2 to the first orbital and 8 to the second
 - b) What particle(s) are found in the center of the atom?
Protons and neutrons
3. Play until you discover which **particle(s)** determine(s) the name of the **element** you build. What did you discover?
Protons - adding them moves from left to right on the Periodic Table
4. What is the **name** of the following atoms?
 - a) An atom with 3 protons and 4 neutrons: **Lithium**
 - b) An atom with 2 protons and 4 neutrons: **Helium**
 - c) An atom with 4 protons and 4 neutrons: **Beryllium**
5. Play with the simulation to discover which particles affect the **charge** of an atom or ion.
 - a) Fill in the blanks below to show your results:

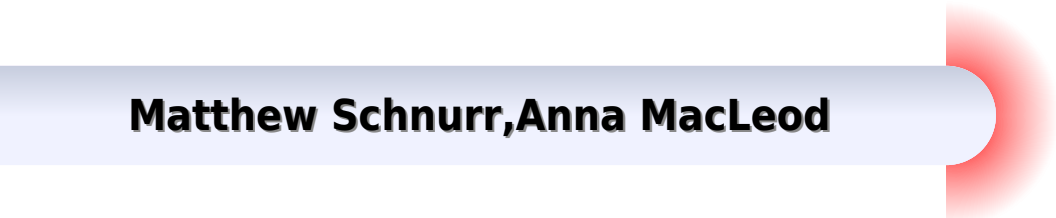
Neutral atoms have the same number of protons and electrons.

Positive ions have more protons than electrons.

Negative ions have fewer protons than electrons.
 - b) Develop a relationship (in the form of a single sentence or equation) that can predict the charge based on the number and types of particle.
The charge of an atom is determined by the ratio of protons to electrons.

Build An Atom Phet Lab Worksheet Answers

Matthew Schnurr, Anna MacLeod



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Table of Contents Build An Atom Phet Lab Worksheet Answers

1. Understanding the eBook Build An Atom Phet Lab Worksheet Answers
 - The Rise of Digital Reading Build An Atom Phet Lab Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Build An Atom Phet Lab Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Build An Atom Phet Lab Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Build An Atom Phet Lab Worksheet Answers
 - Personalized Recommendations
 - Build An Atom Phet Lab Worksheet Answers User Reviews and Ratings
 - Build An Atom Phet Lab Worksheet Answers and Bestseller Lists
5. Accessing Build An Atom Phet Lab Worksheet Answers Free and Paid eBooks
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 - Build An Atom Phet Lab Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Build An Atom Phet Lab Worksheet Answers
 - Highlighting and Note-Taking Build An Atom Phet Lab Worksheet Answers
 - Interactive Elements Build An Atom Phet Lab Worksheet Answers
- 8. Staying Engaged with Build An Atom Phet Lab Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Build An Atom Phet Lab Worksheet Answers
- 9. Balancing eBooks and Physical Books Build An Atom Phet Lab Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Build An Atom Phet Lab Worksheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Build An Atom Phet Lab Worksheet Answers
 - Setting Reading Goals Build An Atom Phet Lab Worksheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Build An Atom Phet Lab Worksheet Answers
 - Fact-Checking eBook Content of Build An Atom Phet Lab Worksheet Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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