

Data and Calculations: The Atomic Spectrum of Hydrogen

A. Calculation of the Energy Levels of the Hydrogen Atom

Energies are to be calculated from Equation 6 for the 10 lowest energy states.

Table 11.2

Quantum Number, n	Energy, E_n , in kJ/mole	Quantum Number, n	Energy, E_n , in kJ/mole
1	-2.16×10^{-21}	6	-6.01×10^{-23}
2	-5.41×10^{-22}	7	4.42×10^{-23}
3	-2.4×10^{-22}	8	2.67×10^{-23}
4	-1.35×10^{-22}	9	2.67×10^{-23}
5	8.00 -8.66×10^{-23}	10	2.16×10^{-23}

B. Calculation of Wavelengths in the Spectrum of the H Atom

In the upper half of each box write ΔE , the difference in energy in kJ/mole between E_{n_u} and E_{n_l} . In the lower half of the box, write λ in nm associated with that value of ΔE .

Table 11.3

n_u	6	5	4	3	2	1
1	2.1×10^{-21} 94.2 nm	2.07×10^{-21} 95.6 nm	2.02×10^{-21} 98.01 nm	1.92×10^{-21} 105.1 nm	3.24×10^{-22} 611 nm	
2	4.8×10^{-22} 412.5 nm	4.52×10^{-22} 437.7 nm	4.05×10^{-22} 488.8 nm	3×10^{-22} 660 nm		
3	1.8×10^{-22} 1100 nm	1.5×10^{-22} 1320 nm	1.05×10^{-22} 1885.7 nm			
4	7.5×10^{-23} 2440 nm	4.8×10^{-23} 4125 nm				
5	2.6×10^{-23} 7615.3 nm					

$$\Delta E = E_{n_u} - E_{n_l}$$

$$\lambda (\text{nm}) = \frac{1.19627 \times 10^9}{\Delta E}$$

(continued on following page)

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