# AP Chemistry: Acid-Base Worksheet

AP Chemistry Acid-Base WS 0809.doc

Name: \_\_\_\_\_\_ Date: \_\_\_\_\_ Per: \_\_\_\_

 In the following equations, label each species as an acid or a base. Show the conjugate acid-base pairs.

a.  $HCO_3^-(eq) + HF_{(eq)} \leftrightarrow H_2CO_{3(eq)} + F_{(eq)}$ b.  $HCO_3^-(eq) + OH^-_{(eq)} \leftrightarrow CO_3^{-2}_-(eq) + H_3O_{(f)}$ 

2. In the following reactions, identify the Lewis acid and the Lewis base.

a.  $Ag^* + 2NH_2 \leftrightarrow Ag(NH_3)_2$ 

b. B(OH)<sub>3</sub> + H<sub>2</sub>O ↔ B(OH)<sub>4</sub> + H\*

For the following reaction, decide which species (reactants or products) are favored at the completion of the reaction.

 $SO_{1}^{2}(aq) + HCN(aq) \leftrightarrow HSO_{1}(aq) + CN'(aq)$ 

4. Which member of each of the following pairs is the stronger acid?

a. NH<sub>a</sub>, PH<sub>a</sub>

e. HSO, , HSeO,

b. HI, H.Te

c. HSO<sub>3</sub> , H<sub>3</sub>SO<sub>3</sub>

d. H-AsO., H-AsO.

5. Calculate the concentrations of hydronium ion and hydroxide ion at 25°C in

a. 0.15M HNO,

b. 0.010M Ca(OH),

- 6. A sample of orange juice has a hydronium-ion concentration of 2.9x10<sup>-1</sup>M. What is the pH? Is the solution acidic?
- 7. The pH of human arterial blood is 7.40. What is the hydronium ion concentration?

# **Ap Worksheet 06d Acids And Bases Answer**

**Roman Wölfel** 

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