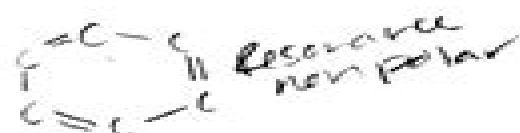


D
D
A
A
D
B
A
D
D
B
most polar
E
A
B
D
E
B
A
A
C
n



Linear 2.0 sp
2.3 sp3d



hydrogen bonding



π - double triple bonds

$$P_{\text{H}_2\text{O}} = \left(\frac{1}{10}\right)(780)$$

dipole-dipole vs London

Boiling point ... = ... (4) B

(6) A
(7) A
(8) C



All 1:1
50 chance
will be 0.2

$$v_{\text{rms}} = \sqrt{\frac{3RT}{M}}$$

Constant

(9) A the fastest - most gone

Therm

(1) B

$$\text{Bond Enthalpy} = \sum \text{broken} - \sum \text{formed}$$

$$= (150 + 3(248)) - (6(216))$$

(2) C

S → L EQ

(3) D

S → L → g entropy

(4) C

$$-889.1 = (-393.3 + x(25.5)) - x$$

$$-889.1 = -965.1 - x$$

$$x = -76$$

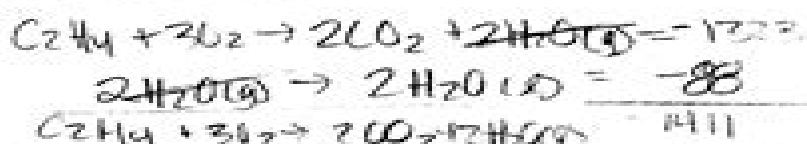
(5) E

S → L entropy, endo, spont, + + -

(6) A

(7) B

(8) E



Phase

(1) E

(2) A

(3) E (Ag is too small to filter)

(4) B

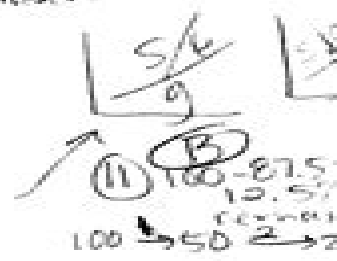
(6) B

(7) B

(8) C

(9) A

(10) A Not in water 12



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Adrian Dingle



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