#### Chemistry Unit 3 Test Review

#### 5.1 Mole and Avogadro's Constant

A particle is very small and can be:

- A single atom (Ca)
- A molecule with 2 or more atoms (molecular compound H2O)
- An ion with a positive or negative charge (CO<sub>1</sub>2-)
- A formula unit with two or more ions (ionic compound NaCl)

#### The Mole

- · Chemists use the mole to represent a specific number of particles
- The number of particles it represents is called the Avogadro constant (6.02 × 10<sup>29</sup> particles/mol)
- The unit for mole is mol

#### Definitions

- Mole: One mole of any entity is 6.02 × 10<sup>25</sup> of the particles (atom, molecule, ion formula unit); the symbol is n; the unit is mol
- Avogadro's constant: The number of particles in one mole, 6.02 × 10<sup>28</sup>; the symbol is Na; the unit is particles/mol
- 3. Particles:
- Atom or element (Eg. Zn)
- Molecule or covalent compound (Eg. H2O)
- lons or monatomic/polyatomic ions (Eg. Zn+2 and CO<sub>1</sub>2-)
- Formula unit or ionic compound (Eg. LiCl)

#### 5.2 Molar Mass

#### Basics: The Mass and The Mole

- The mass of carbon-12 is 12 atomic mass unit (12 u). Every other element has mass relative to one atom of carbon-12
- 1 mole of carbon-12 = 6.02 × 10<sup>23</sup> atoms of C-12 = 12g
- The average atomic mass of any element, in atomic mass units, is numerically
  equivalent to the mass of one mole of atoms (6.02 × 10<sup>28</sup> atoms) of that element

Motar Mass: The mass of one mole of its particle – atoms, molecules, ions, or formula units, in grams per mole (g/mol)

#### 6.1 Percentage Composition

#### Law of Definite Proportions:

- · Elements in a chemical compound are always present in the same proportions by mass
- Law of constant composition: all molecules or formula units of a compound will have the same chemical composition

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**Liying Dong** 

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- 1. Understanding the eBook Apex Chemistry Unit 3 Test
  - The Rise of Digital Reading Apex Chemistry Unit 3 Test
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Apex Chemistry Unit 3 Test
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Apex Chemistry Unit 3 Test
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Apex Chemistry Unit 3 Test
  - Personalized Recommendations
  - Apex Chemistry Unit 3 Test User Reviews and Ratings
  - Apex Chemistry Unit 3 Test and Bestseller Lists

- 5. Accessing Apex Chemistry Unit 3 Test Free and Paid eBooks
  - o Apex Chemistry Unit 3 Test Public Domain eBooks
  - Apex Chemistry Unit 3 Test eBook Subscription Services
  - Apex Chemistry Unit 3 Test Budget-Friendly Options
- 6. Navigating Apex Chemistry Unit 3 Test eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Apex Chemistry Unit 3 Test Compatibility with Devices
  - Apex Chemistry Unit 3 Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Apex Chemistry Unit 3 Test
  - Highlighting and Note-Taking Apex Chemistry Unit 3 Test
  - Interactive Elements Apex Chemistry Unit 3 Test
- 8. Staying Engaged with Apex Chemistry Unit 3 Test
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - $\circ\,$  Following Authors and Publishers Apex Chemistry Unit 3 Test
- 9. Balancing eBooks and Physical Books Apex Chemistry Unit 3 Test
  - Benefits of a Digital Library
  - $\circ\,$  Creating a Diverse Reading Collection Apex Chemistry Unit 3 Test
- 10. Overcoming Reading Challenges
  - o Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Apex Chemistry Unit 3 Test
  - Setting Reading Goals Apex Chemistry Unit 3 Test
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Apex Chemistry Unit 3 Test
  - Fact-Checking eBook Content of Apex Chemistry Unit 3 Test
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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