

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

Crystalline Solids

- Have a regular, repeating arrangement of particles (atoms, ions, or molecules).
- Have a sharp melting point.
- Have a definite shape and volume.
- Have a high density.
- Have a high heat of fusion.
- Have a high heat of vaporization.
- Have a high heat of sublimation.
- Have a high boiling point.
- Have a high boiling point.
- Have a high boiling point.

Amorphous Solids

- Do not have a regular, repeating arrangement of particles.
- Do not have a sharp melting point.
- Do not have a definite shape and volume.
- Do not have a high density.
- Do not have a high heat of fusion.
- Do not have a high heat of vaporization.
- Do not have a high heat of sublimation.
- Do not have a high boiling point.
- Do not have a high boiling point.
- Do not have a high boiling point.

Periodic Table & Ionization Energies

Ionization energy values are provided for elements 1 through 118, showing trends across the periodic table.

Properties of the Periodic Table

- Groups:** Elements in the same group have similar chemical properties.
- Periods:** Elements in the same period have the same number of electron shells.
- Metals:** Elements on the left side of the periodic table are metals.
- Nonmetals:** Elements on the right side of the periodic table are nonmetals.
- Transition Metals:** Elements in the d-block are transition metals.
- Lanthanides and Actinides:** Elements in the f-block are lanthanides and actinides.

Electronegativity Table

Electronegativity increases from left to right and from bottom to top.

Electronegativity values are provided for elements 1 through 118, showing trends across the periodic table.

Types of Chemical Calculations

Stoichiometry

- Mass calculations
- Molar calculations
- Limiting reagent calculations
- Percent yield calculations
- Gas calculations
- Solution calculations
- Acid-base calculations
- Redox calculations
- Thermodynamics calculations
- Equilibrium calculations
- Kinetics calculations
- Electrochemistry calculations
- Atomic structure calculations
- Molecular structure calculations
- Intermolecular forces calculations
- Phase changes calculations
- Colligative properties calculations
- Crystal field theory calculations
- Magnetism calculations
- Optics calculations
- Acoustics calculations
- Thermodynamics calculations
- Equilibrium calculations
- Kinetics calculations
- Electrochemistry calculations
- Atomic structure calculations
- Molecular structure calculations
- Intermolecular forces calculations
- Phase changes calculations
- Colligative properties calculations
- Crystal field theory calculations
- Magnetism calculations
- Optics calculations
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Thermodynamics

First Law of Thermodynamics

$$\Delta E = q + w$$

Second Law of Thermodynamics

$$\Delta S_{univ} > 0$$

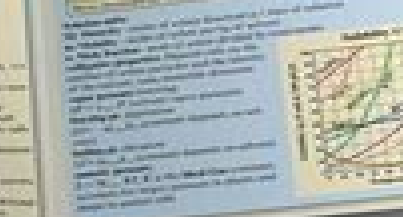
Third Law of Thermodynamics

$$\Delta S_{univ} = 0$$

Equilibrium & Solubility

Equilibrium Constants

Reaction	K_c	K_p	K_a	K_b	K_{sp}	$K_{f/c}$
$2H_2(g) + O_2(g) \rightleftharpoons 2H_2O(g)$	5.0	1.0				
$PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$	0.1	0.1				
$AgCl(s) \rightleftharpoons Ag^+(aq) + Cl^-(aq)$					1.8 x 10 ⁻¹⁰	
$H_2O(l) \rightleftharpoons H^+(aq) + OH^-(aq)$			1.0 x 10 ⁻¹⁴			



Acids & Bases

Arrhenius Definition

- Acids: H^+ ions
- Bases: OH^- ions

Brønsted-Lowry Definition

- Acids: H^+ donors
- Bases: H^+ acceptors

Lewis Definition

- Acids: e^- acceptors
- Bases: e^- donors

Redox Reactions

Oxidation

- Loss of electrons
- Increase in oxidation number

Reduction

- Gain of electrons
- Decrease in oxidation number

Half-Reactions

$$Zn \rightarrow Zn^{2+} + 2e^-$$

$$Cu^{2+} + 2e^- \rightarrow Cu$$

Cell Notation

$$Zn | Zn^{2+} || Cu^{2+} | Cu$$

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