

trial	$[A_2]$	$[B]$	initial rate ($M s^{-1}$)
1	0.10	0.50	2.5×10^{-3}
2	0.20	0.50	5.0×10^{-3}
3	0.30	0.05	5.0×10^{-3}
4	0.30	0.10	1.0×10^{-2}

An experiment was conducted to determine the rate law for the reaction $A_2(g) + B(g) \rightarrow A_2B(g)$. The table above shows the data collected. Based on the data in the table, which statement is correct?

- A** Since the rate law can be expressed as $\text{rate} = k[A_2]^2$, tripling the concentration of A_2 will cause a 9-fold increase in the rate of the reaction.
- B** Since the rate law can be expressed as $\text{rate} = k[A_2][B]$, doubling the concentrations of A_2 and B will quadruple the rate of the reaction.
- C** Since the rate law can be expressed as $\text{rate} = k[A_2]^2[B]$, tripling the concentration of A_2 while keeping the concentration of B constant will triple the rate of the reaction.
- D** Since the rate law can be expressed as $\text{rate} = k[A_2][B]^2$, doubling the concentration of B while keeping the concentration of A_2 constant will double the rate of the reaction.

Answers To Chemfile Experiment A2

RS Peters



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