# REDOX REACTIONS

Name \_\_\_\_

Oxidizing agent

For the equations below, identify the substance oxidized, the substance reduced, the oxidizing agent, the reducing agent, and write the oxidation and reduction half reactions.

1. 2H, + 0, - 2H,0 OX: 2H, -> 4H + + 4e Loducing agent RED: 0, + 4e -> 20^2 - oxidizing agent 2. Fe + Zn2 - Fe2 + Zn EX: Fe -> Fe2 + 2e Leducing agent

3. 2Al + 3Fe<sup>12</sup> - 2Al<sup>3</sup> + 3Fe  $OX : \partial Al^5 \longrightarrow \partial Al^{13} + 6e^-$  Ladueing agent

Red:  $3Fe^{+2} + 6e^- \longrightarrow 3Fe$  Oxidizing agent

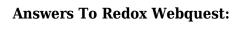
4. Cu + 2AgNO, + CUKNO), + 2Ag

OX: Cu° -> Cu+2 + 2e reducing agent

red: 2Ag+ + 2e -> 2Ag° oxidizing agent

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Yijin Wang



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