

# THE OFFICIAL STUDY GUIDE FOR: "CHEMICAL KINETICS"

**Multiple Choice Section:** This study guide is a compilation of questions from provincial exams since April 1994. I urge you to become intimately familiar with question types. You will notice that questions from one year to another are very similar in their composition. Identification of question types will allow you to be more efficient in answering these questions on the provincial examination. My recommendations for using this study guide are as follows:

- DO ALL THE QUESTIONS** in this booklet. These are actual Provincial Exam questions! Your own provincial exam and unit test will include questions similar to the ones in this booklet!
- RESIST THE URGE TO LOOK AT THE ANSWER KEY** until you have given all the questions in the section your best effort. Don't do one question, then look at the key, then do another and look at the key, and so on. Each time you look at one answer in the study guide, your eye will notice other answers around them, and this will reduce the effectiveness of those questions in helping you to learn.
- LEARN FROM YOUR MISTAKES!** If you get a question wrong, figure out why! If you are having difficulty, talk to your study partner, or maybe phone someone in your Peer Tutoring group. Get together with group members or other students from class and work on these questions together. Explain how you got your answers to tough questions to others. In explaining yourself to someone else, you will learn the material better yourself (try it!) Ask your teacher to explain the questions to you during tutorial or after school. Your goal should be to get 100% on any Chemistry 12 multiple choice test! Learning from your mistakes in this booklet will really help you in your efforts to meet this goal!
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- Check Off the STATUS box on the PRESCRIBED LEARNING OUTCOMES sheet.** I have tried to organize the questions in the identical sequence to which they appear on your Acid Base Prescribed Learning Outcome sheet. By doing this, you can be confident that you know everything you need to know for both the UNIT EXAM and PROVINCIAL EXAM!

## TABLE OF CONTENTS:

|                                        |    |
|----------------------------------------|----|
| INTRODUCTION TO REACTION KINETICS..... | 1  |
| COLLISION THEORY.....                  | 6  |
| REACTION MECHANISMS AND CATALYSTS..... | 17 |
| <b>ANSWERS</b> .....                   | 21 |

## INTRODUCTION

- A2 The rate of a chemical reaction can be expressed in  
 A. grams per mole. B. energy consumed per mole.  
 C. volume of gas per unit time. D. moles formed per litre of solution.
- A2 Consider the following reaction:  
 Under certain conditions, the rate of decomposition of  $\text{NO}_2$  is  $3.2 \times 10^{-3} \text{ mol/s}$ . The rate of formation of  $\text{O}_2$  is:  
 A.  $1.6 \times 10^{-3} \text{ mol/s}$   
 B.  $3.2 \times 10^{-3} \text{ mol/s}$   
 C.  $4.8 \times 10^{-3} \text{ mol/s}$   
 D.  $6.4 \times 10^{-3} \text{ mol/s}$
- A2 Consider the following reaction:  

$$2\text{N}_2\text{O}_{5(g)} \rightarrow 4\text{NO}_{2(g)} + \text{O}_{2(g)}$$
  
 At a certain temperature the rate of decomposition of  $\text{N}_2\text{O}_5$  is  $2.5 \times 10^{-6} \text{ mol/s}$ . The rate of formation of  $\text{NO}_2$  is:  
 A.  $1.0 \times 10^{-6} \text{ mol/s}$  B.  $1.3 \times 10^{-6} \text{ mol/s}$  C.  $2.5 \times 10^{-6} \text{ mol/s}$  D.  $5.0 \times 10^{-6} \text{ mol/s}$
- A2 Consider the following reaction:  

$$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$$
  
 The rate of formation of  $\text{NH}_3$  is  $3.0 \text{ mL/min}$ . The rate of consumption of  $\text{H}_2$  is  
 A.  $1.5 \text{ mL/min}$  B.  $2.0 \text{ mL/min}$  C.  $4.5 \text{ mL/min}$  D.  $9.0 \text{ mL/min}$
- A2 Which of the following units can be used to express the rate of a chemical reaction?  
 A.  $\text{mL/g}$  B.  $\text{mol/L}$  C.  $\text{g/mol}$  D.  $\text{mol/min}$

# Answers To Chemical Kinetics Study Guide

**Clyde Metz**



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1. Understanding the eBook Answers To Chemical Kinetics Study Guide
  - The Rise of Digital Reading Answers To Chemical Kinetics Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Answers To Chemical Kinetics Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answers To Chemical Kinetics Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answers To Chemical Kinetics Study Guide
  - Personalized Recommendations
  - Answers To Chemical Kinetics Study Guide User Reviews and Ratings

- Answers To Chemical Kinetics Study Guide and Bestseller Lists
- 5. Accessing Answers To Chemical Kinetics Study Guide Free and Paid eBooks
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- 8. Staying Engaged with Answers To Chemical Kinetics Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answers To Chemical Kinetics Study Guide
- 9. Balancing eBooks and Physical Books Answers To Chemical Kinetics Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answers To Chemical Kinetics Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Answers To Chemical Kinetics Study Guide
  - Setting Reading Goals Answers To Chemical Kinetics Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers To Chemical Kinetics Study Guide
  - Fact-Checking eBook Content of Answers To Chemical Kinetics Study Guide
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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