

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

Crystalline Solids

- Have a regular, repeating arrangement of particles (atoms, ions, or molecules).
- Have a definite melting point.
- Have a definite heat of fusion.
- Have a definite density.
- Have a definite color and luster.
- Have a definite refractive index.
- Have a definite cleavage.
- Have a definite crystal structure.
- Have a definite crystal habit.
- Have a definite crystal form.
- Have a definite crystal system.
- Have a definite crystal class.
- Have a definite crystal habit.
- Have a definite crystal form.
- Have a definite crystal system.
- Have a definite crystal class.

Amorphous Solids

- Do not have a regular, repeating arrangement of particles.
- Do not have a definite melting point.
- Do not have a definite heat of fusion.
- Do not have a definite density.
- Do not have a definite color and luster.
- Do not have a definite refractive index.
- Do not have a definite cleavage.
- Do not have a definite crystal structure.
- Do not have a definite crystal habit.
- Do not have a definite crystal form.
- Do not have a definite crystal system.
- Do not have a definite crystal class.

Periodic Table & Small Diagrams

Properties of the Periodic Table

- Atomic Number:** The number of protons in the nucleus of an atom.
- Mass Number:** The sum of the number of protons and neutrons in the nucleus of an atom.
- Atomic Weight:** The average mass of the atoms of an element, weighted by their relative abundance.
- Electronegativity:** The ability of an atom to attract electrons towards itself.
- Ionization Energy:** The energy required to remove an electron from an atom.
- Electron Affinity:** The energy released when an electron is added to a neutral atom.
- Atomic Radius:** The distance from the center of the nucleus to the outer edge of the electron cloud.
- Covalent Radius:** The distance between the nuclei of two atoms bonded together.
- Van der Waals Radius:** The distance between the nuclei of two non-bonded atoms.
- Crystal Radius:** The distance from the center of the nucleus to the center of the nearest neighbor in a crystal lattice.
- Hydrated Radius:** The distance from the center of the nucleus to the center of the hydration shell.
- Effective Nuclear Charge:** The net positive charge experienced by an electron in a multi-electron atom.
- Shielding Constant:** A measure of the extent to which the inner electrons shield the outer electrons from the nucleus.
- Slater's Rules:** A set of rules for calculating the shielding constant for an electron in a multi-electron atom.
- Mulliken's Scale:** A scale for measuring electronegativity based on the average of the ionization energy and electron affinity of an element.
- Pauling's Scale:** A scale for measuring electronegativity based on the relative ability of an element to attract electrons in a chemical bond.
- Allred-Rochow's Scale:** A scale for measuring electronegativity based on the ratio of the atomic number to the square of the atomic radius.
- Shannon's Scale:** A scale for measuring electronegativity based on the electrostatic potential of an ion in a crystal lattice.
- Allen's Scale:** A scale for measuring electronegativity based on the average of the ionization energy and electron affinity of an element, weighted by the number of valence electrons.
- Hammett's Scale:** A scale for measuring electronegativity based on the relative rates of reaction of substituted benzenes.
- Hammett's Constant:** A measure of the relative rates of reaction of substituted benzenes.
- Hammett's Equation:
$$\log \frac{k}{k_0} = \rho \sigma$$**

Electronegativity Table

Electronegativity increases diagonally and decreases the other way, and increases across the periodic table.

- Electronegativity increases from left to right across the periodic table.
- Electronegativity increases from bottom to top across the periodic table.
- Electronegativity increases from left to right across the periodic table.
- Electronegativity increases from bottom to top across the periodic table.

Types of Chemical Calculations

Stoichiometry

- Mass calculations
- Molar calculations
- Volume calculations
- Pressure calculations
- Temperature calculations
- Concentration calculations
- Reaction calculations
- Equilibrium calculations
- Thermodynamics calculations
- Kinetics calculations
- Electrochemistry calculations
- Colligative properties calculations
- Phase diagrams calculations
- Crystal structure calculations
- Crystallography calculations
- Crystal field theory calculations
- Molecular orbital theory calculations
- Quantum mechanics calculations
- Spectroscopy calculations
- Chromatography calculations
- Mass spectrometry calculations
- NMR calculations
- IR calculations
- UV-Vis calculations
- X-ray calculations
- Electron microscopy calculations
- Scanning tunneling microscopy calculations
- Atomic force microscopy calculations
- Scanning electron microscopy calculations
- Transmission electron microscopy calculations
- Scanning probe microscopy calculations
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Properties of Gases

Ideal Gas Law

$$PV = nRT$$

Real Gas Law

$$P(V - nb) = nRT$$

Van der Waals Equation

$$\left(P + \frac{a}{V^2}\right)(V - b) = nRT$$

Boyle's Law

$$PV = \text{constant}$$

Charles's Law

$$\frac{V}{T} = \text{constant}$$

Average Kinetic Energy

$$\frac{1}{2}mv^2 = \frac{3}{2}kT$$

RMS Speed

$$v_{\text{rms}} = \sqrt{\frac{3kT}{m}}$$

Diffusion Coefficient

$$D = \frac{1}{6} \lambda v_{\text{rms}}$$

Mean Free Path

$$\lambda = \frac{1}{\sqrt{2} n \sigma}$$

Collision Frequency

$$Z = \sqrt{2} n \sigma v_{\text{rms}}$$

Maxwell-Boltzmann Distribution

$$f(v) = 4\pi n \left(\frac{m}{2\pi kT}\right)^{3/2} e^{-\frac{mv^2}{2kT}}$$

Most Probable Speed

$$v_{\text{mp}} = \sqrt{\frac{2kT}{m}}$$

Average Speed

$$v_{\text{avg}} = \sqrt{\frac{8kT}{\pi m}}$$

RMS Speed

$$v_{\text{rms}} = \sqrt{\frac{3kT}{m}}$$

Properties of Liquids

Surface Tension

$$\gamma = \frac{F}{L}$$

Viscosity

$$\eta = \frac{\tau}{\dot{\gamma}}$$

Heat of Vaporization

$$\Delta H_{\text{vap}} = \frac{RT_b^2}{T_b - T_c}$$

Heat of Fusion

$$\Delta H_{\text{fus}} = \frac{RT_m^2}{T_m - T_c}$$

Heat of Sublimation

$$\Delta H_{\text{sub}} = \frac{RT_s^2}{T_s - T_c}$$

Heat of Vaporization per Mole

$$\Delta H_{\text{vap}}^{\text{mole}} = \frac{RT_b^2}{T_b - T_c}$$

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Heat of Vaporization per Gram

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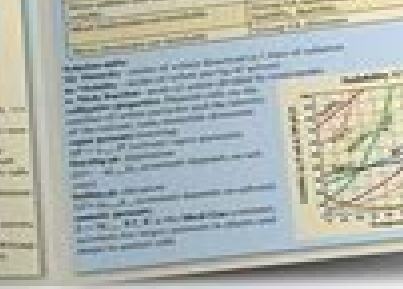
Heat of Vaporization per Atom

$$\Delta H_{\text{vap}}^{\text{atom}} = \frac{RT_b^2}{T_b - T_c}$$

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Annelies Wilder-Smith



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