

Activity 10

Classifying reactions

Can you classify these?

Type of chemical reactions and their descriptions



Introduction: Chemical reactions show how **compounds** and **elements** react with one another. An element is an substance consisting of only one of a kind. Such as the elements of hydrogen, oxygen and fluorine. A compound is a substance made of more than one kind of atoms. Such as water, which is made up of hydrogen.

Question: How are chemical reactions classified?

1. Study these chemical reactions and try to classify each one of them using the chemical equations in the table on the right. Write the following definitions in the right of reaction.

- | | |
|--|-------------------------------|
| ☒ $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
A reaction in which one element reacts with an acid to form a salt and hydrogen gas. | a. Redox |
| ☒ $\text{Ca} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$
A reaction in which a metal reacts with water to form a metal hydroxide and hydrogen gas. | a. Redox |
| ☒ $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
A reaction in which a metal reacts with a salt to form a different salt and a different metal. | c. Single displacement |
| ☒ $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
A reaction in which a metal reacts with a salt to form a different salt and a different metal. | d. Double displacement |
| ☒ $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
A reaction in which a metal reacts with a salt to form a different salt and a different metal. | e. Combustion |

2. Study the reactions in the chemical equations below. Classify each reaction as a synthesis, a decomposition, a single displacement, a double displacement reaction.

- | | |
|---|------------------|
| a. $\text{H}_2\text{O} + \text{CO} \rightarrow \text{H}_2\text{CO}$
Synthesis | Synthesis |
| b. $\text{H}_2\text{O} + \text{CO} \rightarrow \text{H}_2\text{CO}$
Synthesis | Synthesis |
| c. $\text{H}_2\text{O} + \text{CO} \rightarrow \text{H}_2\text{CO}$
Synthesis | Synthesis |
| d. $\text{H}_2\text{O} + \text{CO} \rightarrow \text{H}_2\text{CO}$
Synthesis | Synthesis |
| e. $\text{H}_2\text{O} + \text{CO} \rightarrow \text{H}_2\text{CO}$
Synthesis | Synthesis |

Balancing Equations Gizmo Key

Tobias Bleicker

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