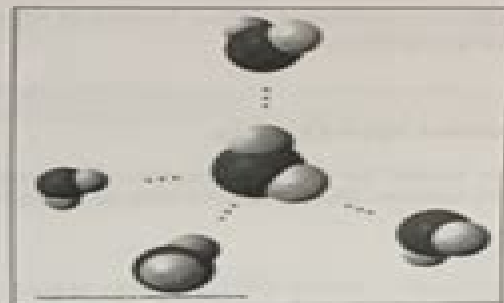


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Chapter 3: Water and the Fitness of the Environment

Concept 3.1 The polarity of water molecules results in hydrogen bonding

1. Study the water molecules at the right. On the central molecule, label oxygen (O) and hydrogen (H).
2. What is a *polar molecule*? Why is water considered polar?
3. Now, add + and - signs to indicate the charged regions of each molecule. Then, indicate the hydrogen bonds.
4. Explain *hydrogen bonding*. How many hydrogen bonds can a single water molecule form?



Concept 3.2 Four emergent properties of water contribute to Earth's fitness for life

Hydrogen bonding accounts for the unique properties of water. Let's look at several.

Cohesion

5. Distinguish between *cohesion* and *adhesion*.
6. What is demonstrated when you see beads of water on a waxed car hood?
7. Which property explains the ability of a water strider to walk on water?

Moderation of Temperature

8. The *calorie* is a unit of heat. Define *calorie*.
9. Water has high *specific heat*. What does this mean? How does water's specific heat compare to alcohol's?
10. Explain how hydrogen bonding contributes to water's high specific heat.

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S Ben Porath



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