

# Antacid Analysis and Titration – Lab Report Assistant

## Exercise 1: Back Titration of Antacid Neutralization

**Data Table 1.** Antacid Neutralization Data

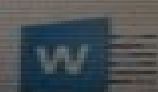
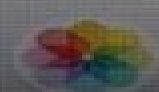
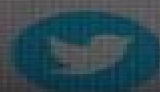
|                                |      |
|--------------------------------|------|
| Mass of Crushed Antacid (g)    |      |
| Concentration of HCl (M)       | 1M   |
| Volume HCl (mL)                | 5 mL |
| Concentration of NaOH (M)      | 1M   |
| Initial NaOH Volume (mL)       |      |
| Final NaOH Volume (mL)         |      |
| Total Volume of NaOH Used (mL) |      |

**Data Table 2.** Experimental Results

|                                                       |                |
|-------------------------------------------------------|----------------|
| HCl available for neutralization (g):                 | 0.3646 (g)     |
| Moles of NaOH required to reach stoichiometric point: | 0.0039 (moles) |
| HCl neutralized by antacid:                           | 0.222 (g)      |
| HCl neutralized per gram of antacid:                  | 0.445 (g)      |

### Questions:

- If an antacid tablet weighed 1.6 grams, how many moles of gastric acid (HCl) would it neutralize? Use the results obtained in **Data Tables 1 and 2** to explain and quantify your answer.
- If you performed this experiment with a titrant of 0.5M NaOH, would you expect your results in **Data Table 2** to change or stay the same? Explain your answer.
- The reaction that occurred when the antacid mixed with the HCl resulted in an additional product (besides a salt and water). Did you see evidence of this product? Describe the experimental evidence you witnessed that supports the formation of the additional product.



# Analysis Of Antacid Manual Answers

**Clemens Wendtner**



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