

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

Basic And Water Chemistry Study Guide Answers

Lingsheng Yao



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Table of Contents Basic And Water Chemistry Study Guide Answers

1. Understanding the eBook Basic And Water Chemistry Study Guide Answers
 - The Rise of Digital Reading Basic And Water Chemistry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Basic And Water Chemistry Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Basic And Water Chemistry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Basic And Water Chemistry Study Guide Answers
 - Personalized Recommendations
 - Basic And Water Chemistry Study Guide Answers User Reviews and Ratings
 - Basic And Water Chemistry Study Guide Answers and Bestseller Lists

5. Accessing Basic And Water Chemistry Study Guide Answers Free and Paid eBooks
 - Basic And Water Chemistry Study Guide Answers Public Domain eBooks
 - Basic And Water Chemistry Study Guide Answers eBook Subscription Services
 - Basic And Water Chemistry Study Guide Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Basic And Water Chemistry Study Guide Answers Compatibility with Devices
 - Basic And Water Chemistry Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Basic And Water Chemistry Study Guide Answers
 - Highlighting and Note-Taking Basic And Water Chemistry Study Guide Answers
 - Interactive Elements Basic And Water Chemistry Study Guide Answers
8. Staying Engaged with Basic And Water Chemistry Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Basic And Water Chemistry Study Guide Answers
9. Balancing eBooks and Physical Books Basic And Water Chemistry Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Basic And Water Chemistry Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Basic And Water Chemistry Study Guide Answers
 - Setting Reading Goals Basic And Water Chemistry Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Basic And Water Chemistry Study Guide Answers
 - Fact-Checking eBook Content of Basic And Water Chemistry Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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