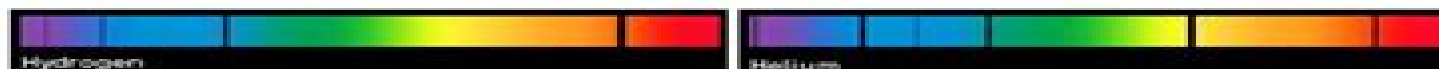


Student Exploration: Bohr Model: Introduction

When light passes through a gas, certain wavelengths of the light are absorbed. The result is a unique **absorption spectrum**. Two examples are shown below.



1. What colors of light are absorbed by hydrogen gas?
violet, blue, green, and red
2. What colors of light are absorbed by helium gas?
violet, blue, green, yellow, and red

Gizmo Warm-up

In 1913, Niels Bohr proposed that the unique spectral lines created by different elements were related to the way electrons were arranged around the nucleus. The *Bohr Model: Introduction* Gizmo explores this connection.

The **laser** shown in the Gizmo can emit **photons**, or particles of light, at a variety of wavelengths. The energy of a photon, measured in **electron volts (eV)**, is inversely proportional to its wavelength. Photons that pass through the gas are detected by the photon detector at the right.



1. With the **Energy (eV)** set to 1 eV, click **Fire**. Did the photon go straight through the gas in the tube, or was it absorbed by the gas?
the photon went straight through the gas tube.
2. Set the **Energy (eV)** to 4 eV, and click **Fire**. What happened this time?
the photon was absorbed by the gas.

Activity A: Absorption spectra

Get the Gizmo ready:

- On the **SIMULATION** pane, select **Lamp**.
- Check that **Gas C** is selected.



Introduction: The smaller the wavelength of a photon, the greater its energy. Humans can see photons with wavelengths between 700 nanometers (red) and 400 nanometers (violet), corresponding to energies between 1.8 and 3.1 eV.

Question: What does the absorption spectrum of an element indicate about its electron configuration?

Bohr Gizmo Answers

RC Schank

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'RC Schank' text.

Bohr Gizmo Answers:

Tribune for Victory and Socialism ,2008

The Atom and the Bohr Theory of Its Structure, an Elementary

Presentation; Hendrik Anthony Kramers,2023-07-18 This book is a clear and concise account of the Bohr model of the atom and its implications Hendrik Kramers explains the principles of atomic structure energy levels and radiation in an accessible way making this book perfect for students and professionals alike This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it This work is in the public domain in the United States of America and possibly other nations Within the United States you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work Scholars believe and we concur that this work is important enough to be preserved reproduced and made generally available to the public We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant **Bohr**

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