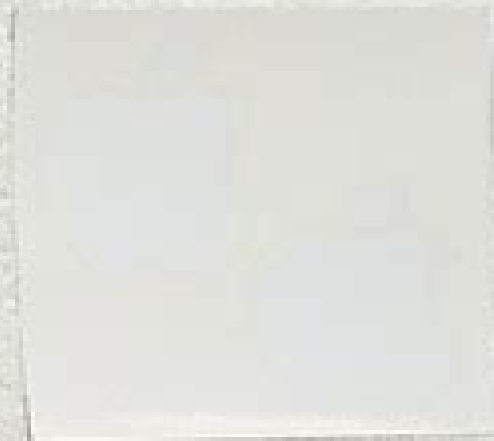




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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained calculated mass of the product expected x100 in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction al o₂ al₂o₃ 4 al 3 o₂ 2 al₂o₃ b determine the theoretical yield of al₂o₃ 160 0 g al x 1 mol al x 2 mol al₂o₃ x 101 964 g al₂o₃ 302 3 g al₂o₃ 26 982 g al 4 mol al 1 mol al₂o₃ c determine the percent yield yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

c is the limiting reagent 4.33 g of H_2 are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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web may 20 2018 learning objectives identify the limiting reactant limiting reagent in a given chemical reaction calculate how much product will be produced from the limiting reactant calculate how much reactant remains when the reaction is complete

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give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3