

7. ELECTROCHEMISTRY

1. ELECTROCHEMISTRY

Electrochemistry is the study of production of electricity from the energy released during a spontaneous chemical reaction and the use of electrical energy to bring about non-spontaneous chemical transformations.

2. ELECTROCHEMICAL CELLS

A spontaneous chemical process is the one which can take place on its own and in such a process the Gibbs's energy of the system decreases. It is this energy that gets converted to electrical energy. The reverse process is also possible in which we can make non-spontaneous processes occur by supplying external energy in the form of electrical energy. These inter conversions are carried out in equipments called Electrochemical Cells.

3. TYPES

Electrochemical Cells are of two types:

3.1 Galvanic Cells

Converts chemical energy into electrical energy

3.2 Electrolytic Cells

Converts electrical energy into chemical energy.

4. GALVANIC CELL

Cell energy is extracted from a spontaneous chemical process or reaction and it is converted to electric current. For example, Daniell Cell is a Galvanic Cell in which Zinc and Copper are used for the redox reaction to take place.



Zn is the reducing agent and Cu^{2+} is the oxidising agent. The half cells are also known as **Electrodes**. The oxidation half is known as **Anode** and the reduction half is called **Cathode**. Electrons flow from anode to cathode in

the external circuit. Anode is assigned **negative polarity** and cathode is assigned **positive polarity**. In Daniell Cell, Zn acts as the anode and Cu acts as the cathode.

5. ELECTROLYTIC CELL

These electrodes are dipped in an electrolytic solution containing cations and anions. On supplying current the ions move towards electrodes of opposite polarity and simultaneous reduction and oxidation takes place.

5.1 Preferential Discharge of Ions

Where there are more than one cation or anion the process of discharge becomes competitive in nature. Discharge of any ion requires energy and in case of several ions being present the discharge of that ion will take place first which requires the energy.

6. ELECTRODE POTENTIAL

It may be defined as the tendency of an element, when it is placed in contact with its own ions to either lose or gain electrons and in turn become positively or negatively charged.

The electrode potential will be named as oxidation or reduction potential depending upon whether oxidation or reduction has taken place.



6.1 Characteristics

(a) Both oxidation and reduction potentials are equal in magnitude but opposite in sign.

(b) It is not a thermodynamic property, so values of E are not additive.

7. STANDARD ELECTRODE POTENTIAL ($E^{\ominus}$)

It may be defined as the electrode potential of an electrode determined relative to standard hydrogen electrode under standard conditions. The standard conditions taken are :

Basic Electrochemistry Manuals

G.P. Simon



Basic Electrochemistry Manuals:

Handbook of Electrochemistry Cynthia G. Zoski, 2007-02-07 Electrochemistry plays a key role in a broad range of research and applied areas including the exploration of new inorganic and organic compounds biochemical and biological systems corrosion energy applications involving fuel cells and solar cells and nanoscale investigations The Handbook of Electrochemistry serves as a source of electrochemical information providing details of experimental considerations representative calculations and illustrations of the possibilities available in electrochemical experimentation The book is divided into five parts Fundamentals Laboratory Practical Techniques Applications and Data The first section covers the fundamentals of electrochemistry which are essential for everyone working in the field presenting an overview of electrochemical conventions terminology fundamental equations and electrochemical cells experiments literature textbooks and specialized books Part 2 focuses on the different laboratory aspects of electrochemistry which is followed by a review of the various electrochemical techniques ranging from classical experiments to scanning electrochemical microscopy electrogenerated chemiluminescence and spectroelectrochemistry Applications of electrochemistry include electrode kinetic determinations unique aspects of metal deposition and electrochemistry in small places and at novel interfaces and these are detailed in Part 4 The remaining three chapters provide useful electrochemical data and information involving electrode potentials diffusion coefficients and methods used in measuring liquid junction potentials serves as a source of electrochemical information includes useful electrochemical data and information involving electrode potentials diffusion coefficients and methods used in measuring liquid junction potentials reviews electrochemical techniques incl scanning electrochemical microscopy electrogenerated chemiluminescence and spectroelectrochemistry **Electrochemical**

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for the monitoring prevention and control of biocorrosion The first part of the book covers the fundamental aspects of microbiology electrochemistry and biofouling of metal surfaces The second half describes biocorrosion assessment in the laboratory and the field the main control and mitigation procedures used practical case studies and laboratory methods and formulations The Manual of Biocorrosion is the book the industrial sector water treatment plants oil refineries etc has been waiting for providing the basics for implementing prevention control and mitigation procedures In addition it covers the latest industry trends with discussions of biocide selection strategies for treating biocorrosion without harming the environment and the latest monitoring programs The academic sector will benefit as well from the up to date information on mechanisms and recent advances in all biocorrosion aspects and technology Research trends such as the application of surface analysis techniques and modern electron microscopy the use of conventional and innovative electrochemical techniques for assessment and microbial inhibition of corrosion are all considered Features 100 illustrations provide you with a visual understanding of the problems and techniques discussed 30 tables give you quick access to data 46 suggested readings provide references on books conference and workshop proceedings and special issues of scientific journals and technical publications specifically devoted to biocorrosion and biofouling 454 reference

Electrochemical Methods: Fundamentals and Applications, 2e Student Solutions Manual Allen J. Bard, Larry R. Faulkner, 2002-01-23 Extensive explanations of problems from the text Student Solutions Manual to accompany Electrochemical Methods Fundamentals and Applications 2nd Edition provides fully worked solutions for the problems presented in the text Extensive in depth explanations walk you step by step through each problem and present alternative approaches and solutions where they exist Graphs and diagrams are included as needed and accessible language facilitates better understanding of the material Fully aligned with the text this manual covers thermodynamics mass transfer impedance spectroelectrochemistry and other related topics and appendices provide detailed mathematical reference and digital simulations

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encouragement to get the job done I have also learned much over the years from Dr S Fisher D R Kunin and Dr I Abrams In one way or another they too have some influence however indirect on this modest effort

Manual of Industrial Corrosion Standards and Control ASTM Committee G-1 on Corrosion of Metals,1973 12th Manual M. Baizer Memorial Symposium on Organic Electrochemistry D. Peters,2016-09-21 Springer Handbook of Electrochemical Energy Cornelia Breitzkopf,Karen Swider-Lyons,2016-12-05 This comprehensive handbook covers all fundamentals of electrochemistry for contemporary applications It provides a rich presentation of related topics of electrochemistry with a clear focus on energy technologies It covers all aspects of electrochemistry starting with theoretical concepts and basic laws of thermodynamics non equilibrium thermodynamics and multiscale modeling It further gathers the basic experimental methods such as potentiometry reference electrodes ion sensitive electrodes voltammetry and amperometry The contents cover subjects related to mass transport the electric double layer ohmic losses and experimentation affecting electrochemical reactions These aspects of electrochemistry are especially examined in view of specific energy technologies including batteries polymer electrolyte and biological fuel cells electrochemical capacitors electrochemical hydrogen production and photoelectrochemistry Organized in six parts the overall complexity of electrochemistry is presented and makes this handbook an authoritative reference and definitive source for advanced students professionals and scientists particularly interested in industrial and energy applications

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