

bizhub C35

Quick Guide



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Bizhub C35 Quick Guide

Colin Banfield, John Walkenbach

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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the ph of human blood involves carbonic acid H_2CO_3 and its conjugate base the bicarbonate ion HCO_3^-

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$$\text{acid} + \text{base} \rightarrow \text{water} + \text{salt}$$
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