#### Data and Calculations: The Atomic Spectrum of Hydrogen

#### A. Calculation of the Energy Levels of the Hydrogen Atom

Energies are to be calculated from Equation 6 for the 10 lowest energy states.

Table 11.2							
Quantum Number, s	Energy, E, in k.Denole	Quantum Number, n	Energy, E, in kJ/mole				
	-2.16×10-21	6	-6.01 VIO-33				
2	-5,41 × 10-17	7	4.45 XIO. 3				
3	-2,4110-22	8	267 x10-23				
ч	-1.35×10-22	9	5 PJ X10-33				
5	-8.66 X10-13	10	2.16 × 10-23				

# B. Calculation of Wavelengths in the Spectrum of the H Atom

In the upper half of each box write  $\Delta E$ , the difference in energy in kJ/mole between  $E_{\infty}$  and  $E_{\infty}$ . In the lower half of the box, write  $\lambda$  in nm associated with that value of  $\Delta E$ .

Table 11	Table 11.3						
Mad	6	5	4	3	2	1	
1	2.1 x10 %	2.07 yie	2.02 10 1	1.92 x10 21	3.24vio		
	94.2 00	95.6 nm	98,0100	103. Inm	611 nm		
2		4.53×10 %					
		437.7 nm					
3	1.8 y10 - x3	1.5x0-2	1.05x10				
	1100 nm	1320 nm	1885.7 An				
4	7.5 XIO-23	4.8 X10-23	e main				
	240nm	4125nm					
	2.6 VIO 3			ΔE =	En - En		
	7615.3 m			λ (nm) =	1.19627 × 10 ΔE		

# **Atomic Spectrum Of Hydrogen Lab Answers**

**Michael Seilmaier** 

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