

## Hess law lab

The objective of the present experiment is to determine  $\Delta H_1$  at constant T and P for the reaction:



To do this, the  $\Delta H$ 's of two reactions will be measured along with a  $\Delta H$  value taken from the literature. Specifically the reactions are:



If the temperature is 25.0 °C and the pressure is 1 atm,

$$\Delta H_4 = 2 \Delta H_f^\circ (\text{H}_2\text{O(l)}) = 2 (-285.6 \text{ kJ/mol}) = -571.6 \text{ kJ/mole}$$

where the enthalpy of formation of liquid water is defined as



Equations (2), (3), and (4) can be manipulated to obtain equation (1). Specifically,

$$\Delta H_1 = 2 \Delta H_2 - 2 \Delta H_3 + \Delta H_4 \quad (5)$$

The  $\Delta H_1$  obtained from (5) will be compared to the accepted value and a percent difference will be reported.

The experiment was performed as follows. For the Mg + HCl reaction, 50.0 mLs of 1M HCl solution was added to two nested Styrofoam cups. A thermometer was placed in a hole in a plastic lid which covered the Styrofoam cup, to minimize heat loss to the air. The starting temperature of the HCl solution was recorded. Approx. 0.12g of Mg was added to the HCl solution, the thermometer/lid was quickly replaced, and the temperature was monitored. The maximum temperature was recorded as the final temperature. The procedure for the MgO + HCl reaction was identical, except that approx. 0.50g of MgO was used instead of Mg.

Fill in the data for Table 2 below with the following data.

Line 1: 0.113

Lines 2 and 3: ignore

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## **Table of Contents Answer Key To Hess Lab Manual 10 E**

1. Understanding the eBook Answer Key To Hess Lab Manual 10 E
  - The Rise of Digital Reading Answer Key To Hess Lab Manual 10 E
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Hess Lab Manual 10 E
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Key To Hess Lab Manual 10 E
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Hess Lab Manual 10 E
  - Personalized Recommendations
  - Answer Key To Hess Lab Manual 10 E User Reviews and Ratings
  - Answer Key To Hess Lab Manual 10 E and Bestseller Lists
5. Accessing Answer Key To Hess Lab Manual 10 E Free and Paid eBooks
  - Answer Key To Hess Lab Manual 10 E Public Domain eBooks
  - Answer Key To Hess Lab Manual 10 E eBook Subscription Services
  - Answer Key To Hess Lab Manual 10 E Budget-Friendly Options
6. Navigating Answer Key To Hess Lab Manual 10 E eBook Formats
  - ePub, PDF, MOBI, and More
  - Answer Key To Hess Lab Manual 10 E Compatibility with Devices
  - Answer Key To Hess Lab Manual 10 E Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answer Key To Hess Lab Manual 10 E
  - Highlighting and Note-Taking Answer Key To Hess Lab Manual 10 E
  - Interactive Elements Answer Key To Hess Lab Manual 10 E

8. Staying Engaged with Answer Key To Hess Lab Manual 10 E
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer Key To Hess Lab Manual 10 E
9. Balancing eBooks and Physical Books Answer Key To Hess Lab Manual 10 E
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer Key To Hess Lab Manual 10 E
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Answer Key To Hess Lab Manual 10 E
  - Setting Reading Goals Answer Key To Hess Lab Manual 10 E
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key To Hess Lab Manual 10 E
  - Fact-Checking eBook Content of Answer Key To Hess Lab Manual 10 E
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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