

IDEAL GAS LAW

Name: _____

Use the Ideal Gas Law below to solve the following problems:

$$PV = nRT \quad \text{where } P = \text{pressure in atmospheres}$$

$$V = \text{volume in liters}$$

$$n = \text{number of moles of gas}$$

$$R = \text{Universal Gas Constant}$$

$$0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}$$

$$T = \text{Kelvin temperature}$$

1. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 liter and exerts a pressure of _____ $[P = 10.1325 \text{ atm} \text{ or } 760 \text{ mmHg}]$

2. A gas sample with 0.100 mole of nitrogen occupies a volume of 0.250 liter and exerts a pressure of _____ $[P = 39.4896 \text{ atm} \text{ or } 2961.72 \text{ mmHg}]$

3. A gas sample with 0.500 mole of CO_2 exerts a pressure of 0.500 atm and occupies a volume of 1.00 liter. The temperature of the gas is _____ $[T = 273.15 \text{ K} \text{ or } 0^\circ\text{C}]$

4. A gas sample with 0.500 mole of CO_2 exerts a pressure of 0.500 atm and occupies a volume of 1.00 liter. The temperature of the gas is _____ $[T = 273.15 \text{ K} \text{ or } 0^\circ\text{C}]$

5. A gas sample with 0.500 mole of CO_2 exerts a pressure of 0.500 atm and occupies a volume of 1.00 liter. The temperature of the gas is _____ $[T = 273.15 \text{ K} \text{ or } 0^\circ\text{C}]$

6. A gas sample with 0.500 mole of CO_2 exerts a pressure of 0.500 atm and occupies a volume of 1.00 liter. The temperature of the gas is _____ $[T = 273.15 \text{ K} \text{ or } 0^\circ\text{C}]$

7. A gas sample with 0.500 mole of CO_2 exerts a pressure of 0.500 atm and occupies a volume of 1.00 liter. The temperature of the gas is _____ $[T = 273.15 \text{ K} \text{ or } 0^\circ\text{C}]$

8. A gas sample with 0.500 mole of CO_2 exerts a pressure of 0.500 atm and occupies a volume of 1.00 liter. The temperature of the gas is _____ $[T = 273.15 \text{ K} \text{ or } 0^\circ\text{C}]$

9. A gas sample with 0.500 mole of CO_2 exerts a pressure of 0.500 atm and occupies a volume of 1.00 liter. The temperature of the gas is _____ $[T = 273.15 \text{ K} \text{ or } 0^\circ\text{C}]$

10. A gas sample with 0.500 mole of CO_2 exerts a pressure of 0.500 atm and occupies a volume of 1.00 liter. The temperature of the gas is _____ $[T = 273.15 \text{ K} \text{ or } 0^\circ\text{C}]$

Answers To Ideal Gas Law Packet

Chao Zhang



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