

50/50

Experiment 9 Prelaboratory Assignment

A Volumetric Analysis

Date 03/30

Desk No. _____

1. a. What is the titrant for this experiment?

NaOH (base), sodium hydroxide

- b. Is the indicator generally added to the titrant or the analyte in a titration?

The indicator, phenolphthalein, is added to the analyte, the acid.
It is colorless in an acid but pink in basic solutions.

2. a. What is the primary standard used in this experiment (name and formula)? Define a primary standard.

A primary standard is an accurate mass of a solid substance which is measured, dissolved in water, and then reacted.

↳ $\text{KHC}_8\text{H}_4\text{O}_4$, potassium hydrogen phthalate

- b. What is the secondary standard used in this experiment (name and formula)? Define a secondary standard.

A secondary standard is the standard remaining solution from after the molar concentration of the titrant is measured.

↳ NaOH, sodium hydroxide

3. Distinguish between a stoichiometric point and an endpoint in an acid-base titration.

reaction is complete
because amount
reactants that have
combined

when the indicator changes
colors due to the change
in pH

4. a. When rinsing a buret after cleaning it with soap and water, should the rinse be dispensed through the buret tip or the top opening of the buret? Explain.

The water and rinse should go through the tip. Solution will dispense through and help keep the concentration of the titrant the same.

- b. Experimental Procedure, Part A.4. In preparing the buret for titration, the final rinse is with the NaOH titrant rather than with deionized water. Explain.

The water can alter the values and the NaOH is hygroscopic.

5. Part A.3. Calculate the mass of $\text{KHC}_8\text{H}_4\text{O}_4$ (molar mass = 204.23 g/mol) that reacts with 15 mL of the 0.15 M NaOH solution prepared in Part A.2. Express this mass $\text{KHC}_8\text{H}_4\text{O}_4$ to the correct number of significant figures and record the calculation on the Report Sheet.

$$15 \text{ mL} = 0.015 \text{ L}$$

$$0.015 \text{ L (0.15 M)} = 0.00225 \text{ mol NaOH}$$

$$0.00225 \text{ mol NaOH} \cdot \frac{1 \text{ mol KHC}_8\text{H}_4\text{O}_4}{1 \text{ mol NaOH}} = 0.00225 \text{ mol KHC}_8\text{H}_4\text{O}_4$$

$$0.00225 \text{ mol KHC}_8\text{H}_4\text{O}_4 \times \left(\frac{204.23 \text{ g}}{1 \text{ mol}} \right) = \boxed{0.4595 \text{ g}}$$



Answers To Chemfile Experiment A

G Orfield



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Table of Contents Answers To Chemfile Experiment A

1. Understanding the eBook Answers To Chemfile Experiment A
 - The Rise of Digital Reading Answers To Chemfile Experiment A
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers To Chemfile Experiment A
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers To Chemfile Experiment A
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers To Chemfile Experiment A
 - Personalized Recommendations
 - Answers To Chemfile Experiment A User Reviews and Ratings
 - Answers To Chemfile Experiment A and Bestseller Lists
5. Accessing Answers To Chemfile Experiment A Free and Paid eBooks

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- Answers To Chemfile Experiment A eBook Subscription Services
- Answers To Chemfile Experiment A Budget-Friendly Options
- 6. Navigating Answers To Chemfile Experiment A eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers To Chemfile Experiment A Compatibility with Devices
 - Answers To Chemfile Experiment A Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers To Chemfile Experiment A
 - Highlighting and Note-Taking Answers To Chemfile Experiment A
 - Interactive Elements Answers To Chemfile Experiment A
- 8. Staying Engaged with Answers To Chemfile Experiment A
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers To Chemfile Experiment A
- 9. Balancing eBooks and Physical Books Answers To Chemfile Experiment A
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers To Chemfile Experiment A
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers To Chemfile Experiment A
 - Setting Reading Goals Answers To Chemfile Experiment A
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers To Chemfile Experiment A
 - Fact-Checking eBook Content of Answers To Chemfile Experiment A
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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