

Micro Mole Rockets

Hydrogen and Oxygen Mole Ratio

Introduction

"It will free him from the remaining chains, the shackles of gravity which still tie him to his planet." —Wernher von Braun

The combustion reaction of hydrogen and oxygen is used to produce the explosive energy needed to power the space shuttle. The reaction is also being engineered to serve as a source of combustion energy for fuel cells in electric vehicles. What factors determine the explosiveness of the reaction of hydrogen with oxygen? In this lab, we will generate microscale quantities of hydrogen and oxygen and test their explosive nature, first separately, then in mixtures of various proportions. The goal—to find the most “powerful” gas mixture and use it to launch a rocket across the room!

Concepts

- Mole ratio
- Stoichiometry
- Combustion
- Limiting reactants

Background

Hydrogen, the most abundant element in the universe, is a colorless, odorless gas. It is combustible, which means that it burns quite readily. Hydrogen gas is conveniently generated in the lab by the reaction of zinc metal with hydrochloric acid.

Oxygen, the next abundant element on Earth, is also a colorless, odorless gas. Oxygen gas supports combustion, that is, it must be present for combustible materials to burn. Small-scale quantities of oxygen gas are conveniently generated in the lab by the decomposition of hydrogen peroxide. The decomposition reaction of hydrogen peroxide requires a catalyst to initiate the reaction. A variety of different catalysts, including manganese, manganese dioxide, potassium iodide and even yeast, have been used in this reaction. In this lab, yeast will be used to catalyze the decomposition of hydrogen peroxide and generate oxygen gas.

Experiment Overview

The purpose of this microscale experiment is to generate hydrogen and oxygen and to determine the optimum ratio for their combustion reaction to give water. The optimum ratio will be used to calculate the mole ratio for the reaction of hydrogen and oxygen in a balanced chemical equation. The concept of limiting reactants will be used to explain the results obtained with various hydrogen/oxygen gas mixtures.

Pre-Lab Questions

1. Write the balanced chemical equation for the single replacement reaction of zinc and hydrochloric acid to generate hydrogen gas.
2. Write the balanced chemical equation for the yeast-catalyzed decomposition of hydrogen peroxide to generate oxygen gas and water.
Note: Since a catalyst is not really a reactant or product, it is usually written over the arrow.

Materials

Hydrochloric acid, HCl, 3 M, 15 mL
 Hydrogen peroxide, H₂O₂, 3%, 15 mL
 Yeast suspension, 2%, 5 mL
 Zinc, reagent Zn, about 5 g
 Beaker, 250-mL
 Graduated cylinder, 10-mL
 Marker (permanent pen)
 Paper towels
 Pencil sparker (optional)
 Pipets, Beral-type, graduated, 4
 Rubber stoppers, one-hole, to fit test tubes, 4
 Safety matches
 Spatula
 Test tubes, small, 4
 Test tube rack
 Wood splint

Safety Precautions

Hydrochloric acid is toxic by ingestion and inhalation and is corrosive to skin and eyes. Hydrogen peroxide is a skin and eye irritant. Avoid contact of all chemicals with skin and eyes and notify your teacher immediately in the case of a spill. Wear chemical splash goggles and chemical-resistant gloves and apron. Wash hands thoroughly with soap and water before leaving the laboratory.

Answers To Micro Mole Rockets Lab

Huangqi Zhang



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