

solutions, acids, and bases

1. VOCABULARY REVIEW OF SOLUTIONS

Solute - what is **dissolved** in the solvent to form a solution.

Solvent - the substance that **dissolves** the solute, ordinarily a liquid when the solute and the solvent come together to form a **homogeneous mixture**.

Miscible - two mix in any proportions (have to be mixed) - liquids will pour each over

Immiscible - don't mix in any proportions (have to be mixed)

examples of miscible-immiscible-pairs: water-alcohol, oil-water, oil-alcohol, etc. water-immiscible

Solute - aspect of being dissolved in liquid

Insoluble - a substance (tends to be a solid) that **will not dissolve** in a solvent. an example is sand and water

Suspension - looks like a solution when **stirred**, but **settles** a bit of water (suspensions). also it is a mixture - gaseous mixture where the small solid particles of a substance are spread throughout a liquid without settling.

Unsaturated - solutes with solubility chart when the temperature would be enough to make a substance dissolve but at a greater amount of solute solubility **less solute dissolves than can** (can still go more things)

Saturated - solutes with solubility chart when the temperature-solubility relationship is at the line. in other words a solution with solute has **reached** at the maximum possible amount.

Supersaturated - contains **more than the max. amount** of solute that can be dissolved at the temperature. on the solubility chart this would mean the grams of solute would go above the line.

2. SOLUBILITY VS. TEMPERATURE CHART



Above the line - **supersaturated**

At the line - **saturated**

Under the line - **unsaturated**

3. METHODS OF SEPARATION

Filtration - separate particles by **size**

Distillation - separate by **boiling point** - some mix liquids to be separated.

Evaporation - allows one to remove a **liquid from a solid** by **transforming** the liquid to a **gaseous state**.

Centrifugation - separating molecules by their densities by spinning them apart.

Decanting - process of separating insoluble or suspension solutions by **pouring** away the liquid from the solid.

Chromatography - separation by the **rates** at which components of different solutions move through a medium (ex. gel electrophoresis)

Acid And Bases Study Guide

**John Krenos, Joseph Potenza, Loretta
Jones, Lynn Koplitz, Thomas Spence**



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises Acid-Base Equilibria - Quick Chemistry Review Outline and Handout E Staff, Acid Base Equilibria Quick Review Outline and Handout for All Students Learn and review on the go Use Quick Review Chemistry Notes to help you learn or brush up on the subject quickly You can use the review notes as a reference to understand the subject better and improve your grades Easy to remember facts to help you perform better Perfect study notes for all high school and college students 10 Pages Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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