

Acid-Base Equilibrium Homework Part II

Part A

- Calculate the pH and pOH of a solution containing 7.30 grams of HCl in 1.0 L of solution.

$$7.30 \text{ g} \times \frac{1 \text{ mol}}{36.46 \text{ g}} = 0.200 \text{ mol} / 1.0 \text{ L} = 0.20 \text{ M HCl}$$

Because HCl is a strong acid, it completely dissociates, so $[\text{HCl}] = [\text{H}_3\text{O}^+] = 0.20 \text{ M}$
 $\text{pH} = -\log[0.20] = 0.70$
 $\text{pOH} = 14 - \text{pH} = 13.30$
- Calculate the pH of a 0.10 M nitric acid (HNO_3) solution.

Because HNO_3 is a strong acid, the acid completely dissociates and the concentration of the acid is equal to the concentration of H^+ at equilibrium.
 $[\text{H}_3\text{O}^+] = 0.10 \text{ M}$
 $\text{pH} = -\log(0.10) = 1$
- Calculate the pH of a $5.0 \times 10^{-3} \text{ M}$ solution of sodium hydroxide.

Sodium hydroxide is a strong base, which means it completely dissociates, so $[\text{NaOH}] = [\text{OH}^-]$.
 $[\text{OH}^-] = 5.0 \times 10^{-3}$
 $\text{pOH} = -\log(5.0 \times 10^{-3}) = 1.3$
 $\text{pH} = 14 - 1.3 = 12.7 = 13 \text{ (2 sig figs)}$

Part B

- Decide whether the statements are true about strong acids, weak acids, both, or neither.
 - Always dissociates leaving behind at least some hydronium ion **both**
 - The initial concentration of the acid is equivalent to the final concentration of hydronium ions **strong only**
 - The pH is easily calculated **strong only**
 - The pH must be calculated by the ICE method **weak only**
 - The acid completely dissociates in water **strong only**
 - The equilibrium concentration of undissociated acid is approximately 0 **strong only**
 - K_a is very small **weak only**
- Write the K_a expression for the following acids:
 - $\text{HC}_2\text{H}_3\text{O}_2$

$$\text{HC}_2\text{H}_3\text{O}_2 (\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{C}_2\text{H}_3\text{O}_2^- (\text{aq}) + \text{H}_3\text{O}^+ (\text{aq})$$

$$K_a = \frac{[\text{C}_2\text{H}_3\text{O}_2^-][\text{H}_3\text{O}^+]}{[\text{HC}_2\text{H}_3\text{O}_2]}$$
 - H_2CO_3

$$\text{H}_2\text{CO}_3 (\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{HCO}_3^- (\text{aq}) + \text{H}_3\text{O}^+ (\text{aq})$$

$$K_a = \frac{[\text{HCO}_3^-][\text{H}_3\text{O}^+]}{[\text{H}_2\text{CO}_3]}$$

Acid Base Homework Answers

Morris Hein, Susan Arena, Cary Willard

Acid Base Homework Answers:

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