

Chapter 1 – Matter and Energy

- 1.1 (a) mass; (b) chemical property; (c) mixture; (d) element; (e) energy; (f) physical property; (g) liquid; (h) density; (i) homogeneous mixture; (j) solid state
- 1.2 (a) atom; (b) chemical change; (c) matter; (d) compound; (e) molecule; (f) physical change; (g) gas; (h) potential energy; (i) hypothesis; (j) kinetic energy
- 1.3 When converting to scientific notation, count the number of places you need to move the decimal point. Zeros to the left of the number are always dropped. For example, the number 0.002030 becomes 2.030×10^{-4} and the zeros to the left of 2030 are dropped. The zero to the right is only kept if it is significant (covered later in this chapter). If the decimal point moves right, the exponent decreases. If the decimal moves left the exponent increases.
(a) 2.95×10^3 ; (b) 8.2×10^{-3} ; (c) 6.5×10^2 ; (d) 1.00×10^{-3}
- 1.4 When converting to scientific notation, count the number of places you need to move the decimal point. Zeros to the left of the number are always dropped. For example, the number 0.002030 becomes 2.030×10^{-4} and the zeros to the left of 2030 are dropped. The zero to the right is only kept if it is significant (covered later in this chapter). If the decimal point moves right, the exponent decreases. If the decimal moves left the exponent increases.
(a) 1.0×10^{-3} ; (b) 4.5×10^3 ; (c) 9.01×10^3 ; (d) 7.9×10^{-4}
- 1.5 When converting from scientific notation to standard notation you may need to add place-holder zeros so that the magnitude of the number is correct. For example, to get 1.86×10^{-5} into standard notation, you need to increase the power by five, so the decimal moves to the left. In addition, you'll need four placeholder zeros after the decimal to show the magnitude of the number, 0.0000186.
(a) 0.0000186; (b) 10,000,000; (c) 453,000; (d) 0.0061
- 1.6 When converting from scientific notation to standard notation you may need to add place-holder zeros so that the magnitude of the number is correct. For example, to get 1.86×10^{-5} into standard notation, you need to increase the power by five, so the decimal moves to the left. In addition, you'll need four placeholder zeros after the decimal to show the magnitude of the number, 0.0000186.
(a) 8200; (b) 0.000002025; (c) 0.07; (d) 300,000,000
- 1.7 (a) 6.2×10^3 ; (b) 3.5×10^3 ; (c) 2.9×10^{-3} ; (d) 2.5×10^{-3} ; (e) 8.20×10^2 ; (f) 1.6×10^{-4}
- 1.8 (a) 2.0×10^5 ; (b) 1.5×10^{11} ; (c) 3.0×10^{-10} ; (d) 8.5×10^{-5} ; (e) 8.56×10^2 ; (f) 1.26×10^2
- 1.9 Nonzero digits and zeros between nonzero digits are significant. Zeros at the end of a number and to the right of the decimal are significant. Zeros to the left of the first nonzero digit and in exponentials (i.e., $\times 10^3$) are not significant. The number 0.0950 has three significant digits. The digits "950" are all significant because (1) the 9 and 5 are nonzero, and (2) the zero is significant because it is at the end of the number and to the right of the decimal.
(a) 3; (b) 2; (c) 4; (d) 2; (e) 3
- 1.10 Nonzero digits and zeros between nonzero digits are significant. Zeros at the end of a number and to the right of the decimal are significant. Zeros to the left of the first nonzero digit and in exponentials (i.e., $\times 10^3$) are not significant. The number 0.04350 has four significant digits. The digits "4350" are all significant because (1) the 4, 3, and 5 are nonzero, and (2) the zero is significant because it is at the end of the number and to the right of the decimal.
(a) 3; (b) 4; (c) 4; (d) 4; (e) 2

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Ralph H. Petrucci, Robert K. Wismer



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