

Chemistry Notes

Oxidation:

- Occurs when a species loses an electron
- Reduction: occurs when a species gains electrons
- Redox reactions: transfers of electrons
- Half reaction: shows the electrons lost or gained during the redox reactions

Oxidation (reducing agent): the electrons (e^-) are written on the product side. (right side of the arrow)

Reduction (oxidizing agent): the electrons (e^-) are written on the reactant side

Oxidation = reducing agent $\longrightarrow e^-$ (product side)

- Oxidation is known as a chemical process.
- Examples of oxidation are like iron rusting and wood rotting.

Summary of oxidation is basically the loss of an electron during a reaction by an atom, molecule or, ion. Oxidation occurs during the increase of a molecule, atom or an ion oxidation state is increased. And the opposite is reduction when it happens there is a decrease in oxidation state and an increase in electrons.

When redox happens it means oxidation and reduction can happen together. When oxidation in a metal by oxygen gas could possibly be explained by the metal losing electrons to form cations with oxygens molecules gaining electrons to form the oxygen atoms.

As Cambridge Chemisrty Notes

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