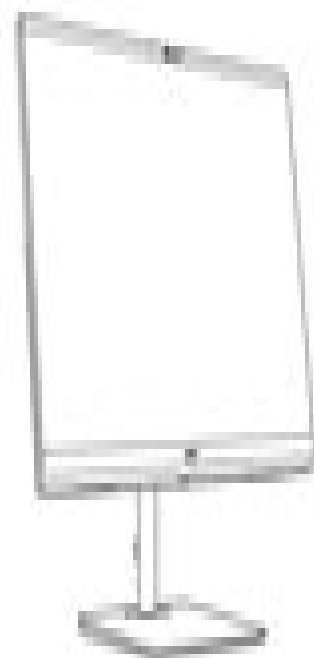




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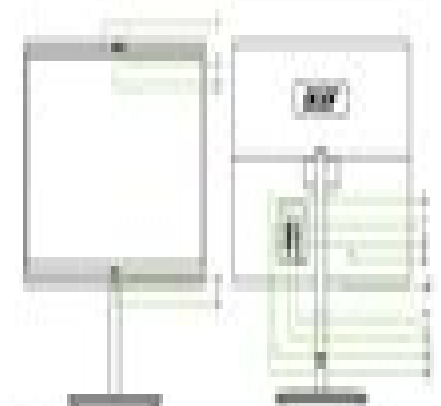
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 $1\text{H}_3\text{PO}_4 + 3\text{NaBr} \rightarrow 1\text{Na}_3\text{PO}_4 + 3\text{HBr}$ type of reaction double displacement
 $2\text{Al(OH)}_3 + 3\text{CaSO}_4 \rightarrow 2\text{Al}_2\text{SO}_4 + 3\text{Ca(OH)}_2$ type of reaction double displacement
 $2\text{Fe} + 3\text{MgO} \rightarrow 2\text{FeO} + 3\text{Mg}$ type of reaction single displacement

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 $\text{NaOH} + \text{KNO}_3 \rightarrow \text{NaNO}_3 + \text{KOH}$ type of reaction double displacement
 $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$ type of reaction combustion
 $2\text{Fe} + 6\text{NaBr} \rightarrow 2\text{FeBr}_3 + 6\text{Na}$ type of reaction single displacement
 $\text{MgSO}_4 + 2\text{NH}_4\text{OH} \rightarrow \text{Mg(OH)}_2 + (\text{NH}_4)_2\text{SO}_4$ type of reaction double displacement
 $\text{C}_4\text{H}_8 + 6\text{O}_2 \rightarrow 4\text{CO}_2 + 4\text{H}_2\text{O}$ type of reaction combustion
 $\text{HCl} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaCl}$ type of reaction double displacement

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 $1\text{H}_3\text{PO}_4 + 3\text{NaBr} \rightarrow 1\text{Na}_3\text{PO}_4 + 3\text{HBr}$ type of reaction double displacement
 $2\text{Al(OH)}_3 + 3\text{CaSO}_4 \rightarrow 2\text{Al}_2\text{SO}_4 + 3\text{Ca(OH)}_2$ type of reaction double displacement
 $2\text{Fe} + 3\text{MgO} \rightarrow 2\text{FeO} + 3\text{Mg}$ type of reaction single displacement

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2 $\text{Ca(OH)}_2 + \text{Al}_2(\text{SO}_4)_3 \rightarrow \text{CaSO}_4 + \text{Al(OH)}_3$ type of reaction 3 $\text{Mg} + \text{Fe}_2\text{O}_3 \rightarrow \text{Fe} + \text{MgO}$ type of reaction 4 $\text{C}_2\text{H}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ type of reaction 5
 $\text{PbSO}_4 + \text{PbSO}_3 + \text{O}_2 \rightarrow$ type of reaction 6 NH_3 12