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Student Exploration: Bohr Model of Hydrogen

Vocabulary: absorption spectrum, Bohr model, electron volt, emission spectrum, energy level, ionization energy, laser, orbital, photon

[Note to teachers and students: This Gizmo was designed as a follow-up to the Bohr Model: Introduction Gizmo. We recommend doing that activity before trying this one.]

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- What happens to an electron when it absorbs a tiny packet of light called a **photon**?
the electron absorbs the energy from it, causing a higher energy level
- What happens to an electron when it emits a photon?
the electrons lose energy, causing a lower energy level

Gizmo Warm-up

When white light is passed through hydrogen gas and separated by a prism, some wavelengths of light are absorbed by the hydrogen atoms in the gas. This causes black bars to appear in the **absorption spectrum** of hydrogen. An **emission spectrum** is created when hydrogen gas emits light.

The **Bohr model** of the atom was inspired by the spectrum produced by hydrogen gas. The link between light and atomic structure is illustrated by the *Bohr Model of Hydrogen* Gizmo. The Gizmo shows a **laser** pointed at a container of hydrogen gas. After passing through the gas, light from the laser goes through a prism and is detected on a screen.

- With the **Laser energy** set to 7.0 eV, click **Play** (). Observe the blue electron on the ORBITALS pane. Do 7.0-eV photons have any effect on the electron? No, it doesn't, the electron stays at lvl 1.
- Click **Pause** (). Set the **Laser energy** to 12.1 eV and click **Play**. What happens to the electron now? the electron shakes up and down and moves up by 2 orbital (level 3)
- Under **Go to energy level**, select **1**. Notice that a photon is emitted by the electron. What is the energy of the emitted photon? (12.1 eV)

Bohr Model Of Hydrogen Gizmo Answer Sheet

Ricardo Simeoni, Griffith University



Bohr Model Of Hydrogen Gizmo Answer Sheet:

Bohr's Model of Atomic Hydrogen Extended to Include Electron Rotational Kinetic Energy Ricardo Simeoni, Griffith University, 2003

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Table of Contents Bohr Model Of Hydrogen Gizmo Answer Sheet

1. Understanding the eBook Bohr Model Of Hydrogen Gizmo Answer Sheet
 - The Rise of Digital Reading Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Bohr Model Of Hydrogen Gizmo Answer Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Personalized Recommendations
 - Bohr Model Of Hydrogen Gizmo Answer Sheet User Reviews and Ratings
 - Bohr Model Of Hydrogen Gizmo Answer Sheet and Bestseller Lists
5. Accessing Bohr Model Of Hydrogen Gizmo Answer Sheet Free and Paid eBooks
 - Bohr Model Of Hydrogen Gizmo Answer Sheet Public Domain eBooks
 - Bohr Model Of Hydrogen Gizmo Answer Sheet eBook Subscription Services
 - Bohr Model Of Hydrogen Gizmo Answer Sheet Budget-Friendly Options

6. Navigating Bohr Model Of Hydrogen Gizmo Answer Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Bohr Model Of Hydrogen Gizmo Answer Sheet Compatibility with Devices
 - Bohr Model Of Hydrogen Gizmo Answer Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Highlighting and Note-Taking Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Interactive Elements Bohr Model Of Hydrogen Gizmo Answer Sheet
8. Staying Engaged with Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Bohr Model Of Hydrogen Gizmo Answer Sheet
9. Balancing eBooks and Physical Books Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Bohr Model Of Hydrogen Gizmo Answer Sheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Setting Reading Goals Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Fact-Checking eBook Content of Bohr Model Of Hydrogen Gizmo Answer Sheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

-
- Interactive and Gamified eBooks

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