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web correct answer c both a and b q2 the density of 1 m solution of hcl is 1 0585 g ml the molality of the solution is 1 0585 1 0 10 0 0585 correct answer b 1 explanation molality begin array l m left frac 1000 m 1000d mm right end array

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web sep 13 2023 the only difference between molality and molarity is in the denominator molality is measured in kilograms of solvent whereas molarity is measured in liters of solution when studying the properties of solutions related to vapour pressure and temperature changes concentrations expressed in molality are used

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web 1 molarity mv mass molar mass $\times 0.4500 \text{ L} = 0.825 \text{ g}$ 141.9579 g/mol $\times 0.0129 \text{ M} = 2 \text{ molality}$ 0.825 g 141.9579 g/mol 0.00581158 mol 0.00581158 mol 0.4500 kg 0.0129 M 3 mole fraction na 2 hpo 4 0.825 g 141.9579 g/mol 0.00581158 mol h 2 o 450.0 g 18.015 g/mol 24.97918401 mol

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web apr 7 2014 molarity moles of solute liters of solution 8.422 a first convert 250 ml to liters $250/1000 = 0.25$ then calculate molarity 5 moles / 0.25 liters = 20 M 3 c a solution with molarity 2 requires 2 M of NaOH per liter so $4 \times 2.8 \text{ M} = 11.2 \text{ M}$ a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture mmix can be calculated using the following formula $m_{\text{mix}} = \frac{m_1 + m_2 + m_3}{V_1 + V_2 + V_3}$ $1.5 \times 480 + 1.2 \times 520 + 480 = 520 + 720 + 624 = 1000$ 1.344 M q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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