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 l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} : \text{MnO}_4^{-}$
 4 : 1 is

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 stoichiometric coefficient of H_2SO_3 if the above equation is balanced in an acidic medium stuck use a hint report a problem
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 redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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 convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{H}_2\text{O} \rightarrow 2\text{HCN} + 2\text{OH}^{-}$ multiplied by a factor of 2 $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ the final answer

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 being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all
 reactions are written as reduction reactions in half reaction tables

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ 2 H_2NO_2 $\text{C}_6\text{H}_6\text{O}_6$ 2 H_2O_2 NO_2 choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ c H_2NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ d H_2O_2 and $\text{C}_6\text{H}_8\text{O}_6$