

Biochemistry Study Guide

1. Organic molecules
 1. **Macromolecules:** Large complex molecules found only in living organisms.
 2. Define the term organic molecule. **Molecules that ave carbon at the center**
 3. Organic molecules are organized around central carbon skeleton.
2. Hydrocarbons
 1. Carbon atoms have four electron in their valence shell, so they can form up to four covalent bonds with other atoms.
 2. Define the term hydro-carbon. **Organic molecules consisting entirely of hydrogen and carbon**
 3. Why are hydrocarbons hydrophobic and insoluble in water? **Hydrocarbons are non polar due to slightly differing electronegativities between carbon and hydrogen**
 4. Describe the difference between aliphatic hydrocarbons and aromatic hydrocarbons. Be able to recognize the structural formula of each. **Aliphatic contain a conjugated bond system, while aromatic do not**
3. Organic molecules are important for living organisms, because of their complex structure.
 1. Organic molecules can form isomers. Define the term isomer. **A molecule that hase the same molecular formula but in different structural formulas**
 2. Organic molecules can form covalent bonds with functional groups. Define the term functional group. **A group of atoms that are attached to the carbon skeleton of an organic molecule**
 3. Be able to name the seven functional groups listed in Figure 2.27 on p. 53, and to recognize their structure.
 4. Be able to recognize each of the seven functional groups within a larger organic molecule.
 5. Macromolecules can be very large.
 6. Name the four groups of macromolecules. **Carbohydrates, protiens, nucleic acids, and lipids**
 7. Carbohydrates, proteins, and nucleic acids can be very large, because they are **polymers** that are built from monomers.
 8. Monomers are linked together during a **dehydration synthesis**. A water molecule is always a product of a dehydration synthesis.

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