

## Flame Test Lab

Name Lani Simmons Period 8<sup>th</sup> Date 10/27

**Objective:** To recognize that the color of light emitted by a heated element is unique and can be translated into energy that is coming from the movement of electrons within the atoms.

**Background:** The natural position of an electron in an energy level is called its *ground state*. When an atom absorbs energy, it is in its *excited state*. This happens when electrons are heated in a flame or electrocuted. These excited electrons jump to higher energy levels. Eventually, the excited electrons fall back down to their ground state, releasing their extra energy as a packet of light called a *photon*.

The color of light corresponds to a specific amount of energy. This color varies from one substance to another, because the location of electrons within each substance is different. Just as we have unique fingerprints, each element has a unique color of light (amount of energy) that it releases. As a result, the colors we observe when a substance is heated may be used to identify the element.

### Pre-lab Questions:

1. Fill in the blanks with the terms "excited, ground, or light."

When an atom absorbs energy, the electrons move from their ground state to their excited state. Then, the electrons release energy, as they fall back down from their excited state to their ground state. This release of energy is in the form of light.

2. What is a photon? A photon is a tiny particle that comprises waves of electromagnetic radiation. As shown by Maxwell, photons are just electric fields traveling through space. Photons have no charge, no resting mass, and travel at the speed of light.

**Data:** Hint: Each color will only be used once in the data table below.

| Metal Salt         | Circle the Color of the Flame that You Observed: |        |            |        |     |                      |
|--------------------|--------------------------------------------------|--------|------------|--------|-----|----------------------|
| Calcium Chloride   | Yellow/Green                                     | Yellow | Blue-green | Orange | Red | Crimson (Darker Red) |
| Copper Chloride    | Yellow/Green                                     | Yellow | Blue-green | Orange | Red | Crimson (Darker Red) |
| Lithium Chloride   | Yellow/Green                                     | Yellow | Blue-green | Orange | Red | Crimson (Darker Red) |
| Sodium Chloride    | Yellow/Green                                     | Yellow | Blue-green | Orange | Red | Crimson (Darker Red) |
| Barium Chloride    | Yellow/Green                                     | Yellow | Blue-green | Orange | Red | Crimson (Darker Red) |
| Strontium Chloride | Yellow/Green                                     | Yellow | Blue-green | Orange | Red | Crimson (Darker Red) |
| Unknown            | Yellow/Green                                     | Yellow | Blue-green | Orange | Red | Crimson (Darker Red) |

Unknown Number \_\_\_\_\_

### Post-lab Questions:

1. In this lab, you tested six different metal salts. Each metal salt contained one metal atom and one chlorine atom. How do you know that the color observed for each salt came from exciting the *metal* atom and not the *chlorine* atom?

2. Identify your unknown metal atom in the space below. *CHLORIDE* should not be part of your answer!

| Unknown Number | Identification |
|----------------|----------------|
|                |                |

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## **Table of Contents Bettelheim Post Lab Answers**

1. Understanding the eBook Bettelheim Post Lab Answers
  - The Rise of Digital Reading Bettelheim Post Lab Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Bettelheim Post Lab Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Bettelheim Post Lab Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Bettelheim Post Lab Answers
  - Personalized Recommendations
  - Bettelheim Post Lab Answers User Reviews and Ratings
  - Bettelheim Post Lab Answers and Bestseller Lists
5. Accessing Bettelheim Post Lab Answers Free and Paid eBooks

- Bettelheim Post Lab Answers Public Domain eBooks
- Bettelheim Post Lab Answers eBook Subscription Services
- Bettelheim Post Lab Answers Budget-Friendly Options
- 6. Navigating Bettelheim Post Lab Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Bettelheim Post Lab Answers Compatibility with Devices
  - Bettelheim Post Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Bettelheim Post Lab Answers
  - Highlighting and Note-Taking Bettelheim Post Lab Answers
  - Interactive Elements Bettelheim Post Lab Answers
- 8. Staying Engaged with Bettelheim Post Lab Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Bettelheim Post Lab Answers
- 9. Balancing eBooks and Physical Books Bettelheim Post Lab Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Bettelheim Post Lab Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Bettelheim Post Lab Answers
  - Setting Reading Goals Bettelheim Post Lab Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Bettelheim Post Lab Answers
  - Fact-Checking eBook Content of Bettelheim Post Lab Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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