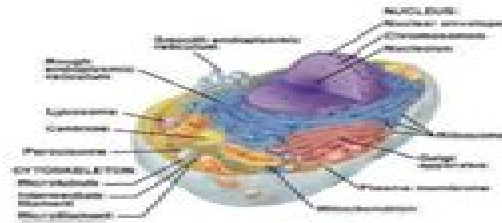
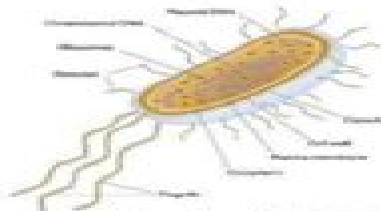


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prokaryotic vs eukaryotic cells

- prokaryotic cells
 - bacteria and archaea
 - have a nucleoid but no nucleus and no true organelles
- eukaryotic cells
 - membrane-bound nucleus and organelles



compartmentalization in eukaryotes

- why are eukaryotic cells so large?
 - compartmentalization: the ability to section off/compartmentalize functions
 - membrane-separated spaces are known as organelles
- why is it important?
 - efficiency: multiple parts of the cell aren't doing the same thing and each part is doing a function they are specialized in
- how did it occur?
 - eukaryotic cells first appeared 2.5 billion years ago
 - internal membranes allow for the compartmentalization of functions
 - endosymbiont theory
 - a prokaryotic cell was engulfed by a larger cell → organelles
 - evidence: mitochondria have their own DNA, and it resembles bacterial DNA; mitochondria undergo binary fission (cell division) to reproduce; mitochondria and chloroplasts have a double membrane indicating one membrane engulfed another
- difference between animal and plant cells
 - lysosomes are not found in plant cells
 - plant cells have cell walls, chloroplasts, and a central vacuole

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web the molar volume of a gas expresses the volume occupied by 1 mole of that respective gas under certain temperature and pressure conditions the most common example is the molar volume of a gas at stp standard temperature and pressure which is equal to 22.4 l for 1 mole of any ideal gas at a temperature equal to 273.15 k and a pressure

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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters

instead of ml we have $v_m = \frac{32 \text{ g mol}^{-1}}{1790 \text{ g l}^{-1}} = 0.0405 \text{ l mol}^{-1}$ nonnumber

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web the molar volume of a gas from avogadro's law the volume is directly proportional to the number of moles the important point to keep in mind is that n can be moles of any gas since one mole always contains 6.02×10^{23} molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as 6.02×10^{23} particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always 6.02×10^{23} but as different particles have different mass their molar mass varies

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web mar 12 2018 you can define the concentration of a gas though and that calculation would be exactly the one you've described placing x moles of a gas in a container of volume v yields a concentration $c = \frac{x}{v}$ with units of $\frac{\text{g}}{\text{pu mol over dm}^3}$

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web the molar volume has the SI unit of cubic metres per mole $\text{m}^3 \text{ mol}^{-1}$ although it is more typical to use the units cubic decimetres per mole $\text{dm}^3 \text{ mol}^{-1}$ for gases and cubic centimetres per mole $\text{cm}^3 \text{ mol}^{-1}$ for liquids and solids

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web zero degrees celsius or 273 K is defined as standard temperature and one atmosphere or 760 mm Hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure STP is 22.4 l

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web at standard temperature and pressure STP the molar volume v_m is the volume occupied by one mole of a chemical element or a chemical compound it can be calculated by dividing the molar mass M by mass density ρ molar gas volume is one mole of any gas at a specific temperature and pressure has a fixed volume molar volume formula

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web after defining the gas phase properties of gases such as temperature pressure and volume are discussed the relationships between these properties are explained at a molecular level and simple models are introduced that allow the various gas laws to be derived from first principles finally the collisional behavior of gases is used to

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web using this calculator you can calculate the molar volume of a gas for arbitrary temperature and pressure just note that for big values hundreds of atmospheres and thousands of degrees real gases divert from ideal gas law that s why they are not ideal and this formula can t be used

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web this volume is known as the molar gas volume V_m the volume of a gas may be calculated from its number of moles using volume of gas moles V_m rearranging this moles of gas $\frac{V}{V_m}$

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web jun 8 2023 $V = 1.00 \text{ mol} \times 0.08206 \text{ latm mol}^{-1} \text{K}^{-1} \times 273 \text{K} = 22.4 \text{ l}$ in other words 1 mole of a gas will occupy 22 4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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web aug 3 2022 volumes of gases and liquids are measured in cubic centimetres cm^3 or cubic decimetres dm^3 or litres l 1 mole of any gas at room temperature and pressure occupies 24 dm^3 $24\,000 \text{ cm}^3$ this value is known as the molar volume

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