

IDEAL GAS LAW

Name: _____

Use the Ideal Gas Law below to solve the following problems:

$$PV = nRT \quad \text{where } P = \text{pressure in atmospheres}$$

$$V = \text{volume in liters}$$

$$n = \text{number of moles of gas}$$

$$R = \text{Universal Gas Constant}$$

$$0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}$$

$$T = \text{Kelvin temperature}$$

pressure: 1 atm
volume: 1 L
n: 1 mol
R: 0.0821

1. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

2. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $V = 1.00 \text{ L}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

3. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

4. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

5. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

6. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

7. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

8. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

9. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

10. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

11. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

12. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

13. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

14. A gas sample of 0.500 mol of nitrogen occupies a volume of 1.00 L at 25°C.
Calculate _____ $P = 10.13 \text{ atm}$ $n = 0.500 \text{ mol}$ $R = 0.0821$ $T = 298 \text{ K}$

Answers To Ideal Gas Law Packet

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