

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

Chapters 17 & 18

Essential Skills/ State Standards:

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CSS: 7d)
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS: 9a)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vocabulary

- | | | | |
|---|---|--|--|
| <ul style="list-style-type: none"> • Condensation • Freezing • Vaporization • Melting • Boiling • Sublimation | <ul style="list-style-type: none"> • Deposition • Heat • Temperature • Thermochemistry • Exothermic • Endothermic | <ul style="list-style-type: none"> • Calorimeter • Joule • Specific heat • $q = m \cdot c \cdot \Delta T$ • $q = m \cdot \Delta H_{\text{vaporization}}$ • $q = m \cdot \Delta H_{\text{fusion}}$ | <ul style="list-style-type: none"> • Heat of reaction • Le Chatelier's principle • Equilibrium • Activation Energy • Catalyst |
|---|---|--|--|

Book Assignments:

Topic	Read	Required Problems
1. Heat, Temperature & Specific Heat	Section 17-1 p. 511-524	Section 17-p. 524 #2, 5 & 6 Chapter 17 Review p. 548 #16 & 17
3. Chemical Equilibrium & Le Chatelier's Principle	Section 12-3 p. 372-375 Section 18-2 p. 562-568	Chapter 18 Review p. 586 #4, 6 & 7
4. Phase Changes Diagrams & Latent Heat of Phase Change	Section 12-3 p. 378-382	Section 12-3 p. 382 #4-8 Chapter 12 Review p. 389 #19, 28, 31 & 37
5. Collision Theory & Activation Energy	Section 17-3 p. 532-534	Section 17-3 p. 537 #2, 3 & 5
6. Reaction Rates & Catalysts	Section 17-4 p. 538-544	Section 17-4 p. 545 #2-4

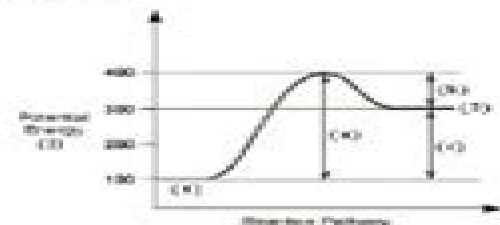
Study Guide (Use notes, worksheets, and labs in addition to this study guide)

- According to the collision theory, what 3 things must happen in order for two molecules to react?
- What is the importance of the activation energy in a chemical reaction?
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:

$$2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$

a. Increase temp:	b. Increase HCl	c. Increase pressure	d. Remove MgCl_2	e. Add H_2
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- Identify the following terms with their correct number/letter from the graphs:

• Reactants-	• Products-
• Activation energy (number/letter & amount in Joules)-	
• Is this reaction exothermic or endothermic?	
- Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do).



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