

Post-Lab Questions:

1. Which of the following compounds would form precipitates in an aqueous solution? NaCN , Na_2CO_3 , FeSO_4 , ZnS , Ca(OH)_2 , CuCO_3

Answer:

- a. NaCN (Sodium Cyanide) - Soluble in water. No precipitate formed.
- b. Na_2CO_3 (Sodium Carbonate) - Soluble in water. No precipitate formed.
- c. FeSO_4 (Iron (II) Sulfate) - Soluble in water. No precipitate formed.
- d. ZnS (Zinc Sulfide) - Insoluble in water. Precipitate formed.
- e. Ca(OH)_2 (Calcium Hydroxide) - Insoluble in water. Precipitate formed.
- f. CuCO_3 (Copper (II) Carbonate) - Insoluble in water. Precipitate formed.

$\therefore$ Therefore, ZnS , Ca(OH)_2 , and CuCO_3 are the compounds that would form precipitates in an aqueous solution.

2. Define an aqueous solution.

Answer: An aqueous solution is a solution in which water is the solvent. This simply means a solution that something has been dissolved in water.

3. The reaction of a sodium hydroxide and sulfuric acid is classified as a _____ reaction.

Answer: The reaction of sodium hydroxide and sulfuric acid is known as a **neutralization reaction**.

4. What would be the net ionic equation for the reaction of sodium hydroxide and sulfuric acid?

Answer: The net ionic equation is $2\text{Na}^+(\text{aq}) + 2\text{OH}^-(\text{aq}) + 2\text{H}^+(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow 2\text{Na}^+(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$

5. What are the spectator ions in the above reaction?

Answer: The spectator ions in the above reaction are Na^+ and SO_4^{2-} . This is because they do not participate in the chemical reaction and remain in the solution in their initial forms.

6. What precipitate is formed from the reaction of barium chloride and silver nitrate?

Answer: From the reaction of barium chloride and silver nitrate, a **white precipitate** is formed.

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J Spring



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